

Name: Ahmed Faizan Date: 12 May Period:

Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units.

Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$\begin{aligned} q &= ? & \Delta t &= 55 - 22 = 33^\circ\text{C} & q &= m s \Delta t \\ m &= 10\text{g} & s &= 0.9 \text{ J/g}^\circ\text{C} & &= 10 \times 0.9 \times 33 \\ & & & & &= \underline{297 \text{ J}} \end{aligned}$$

1. The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did this sample absorb? C for water = 4.18 J/g°C

$$\begin{aligned} \Delta t &= 26.4 - 24.5 = 2 & q &= m s \Delta t \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= 335 \times 4.18 \times 2 \\ m &= 335\text{g} & &= \underline{2800.6 \text{ J}} \end{aligned}$$

2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$\begin{aligned} \Delta t &= 25 - 10 = 15 & q &= m s \Delta t \\ m &= 225\text{g} & &= 225 \times 4.18 \times 15 \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= \underline{14107.5 \text{ J}} \\ 1\text{J} &= 1 \times 10^{-3} \text{ KJ} & \therefore &= \underline{14.1075 \text{ KJ}} \end{aligned}$$

3. To bring 1.0kg of water from 25°C to 99°C takes how much heat input?

$$\begin{aligned} m &= 1\text{kg} = 1000\text{g} & q &= 1000 \times 4.18 \times 74 \\ \Delta t &= 99 - 25 = 74 & q &= \underline{309320 \text{ J}} \\ s &= 4.18 \text{ J/g}^\circ\text{C} & & \end{aligned}$$

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

Imp

$$\begin{aligned} (q_{\text{water}} &= -q_{\text{metal}}) & (C = m s) &\rightarrow s = \frac{C}{m} \\ q_{\text{water}} &= 75 \times 4.18 \times (28.34 - 24) & q &= m s \Delta t \\ &= 1252 & 1252 &= 26 \times s \times (82.25 - 28.34) \\ & & & 1252 = 26 \times 53.91 \times s \\ & & & s = \underline{0.893 \text{ J/g}^\circ\text{C}} \end{aligned}$$

5. A calorimeter has a heat capacity of 1265 J/°C. A reaction causes the temperature of the calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process?

$$\begin{aligned} C &= 1265 \text{ J/}^\circ\text{C} \\ \Delta t &= 25.12 - 22.34 = 2.78 \\ q &= ? \\ q &= C \Delta t \\ &= 1265 \times 2.78 \\ &= \underline{3516.7 \text{ J}} \end{aligned}$$

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12 Rotational Motion and Moment of Inertia Unit 6 Gravitation 13 Gravitation Unit 7 Properties of Bulk Matter 14 Elasticity

15 Pressure of Fluids 16 Viscosity 17 Surface Tension 18 Temperature and Calorimetry 19 Transfer of Heat Unit 8

Thermodynamics 20 First Law of Thermodynamics 21 Second Law of Thermodynamics Unit 9 Behaviour of Perfect Gases and

Kinetic Theory of Gases 22 Behaviour of Perfect Gas and Kinetic Theory of Gases Unit 10 Oscillations and Waves 23

Oscillations 24 Speed of Mechanical Waves and Progressive Waves 25 Superposition of Waves Interference and Beats 26

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Contents Chapter 1 Units and Measurements 3 Chapter 2 Motion in a Straight Line 10 Chapter 3 Motion in a Plane 17

Chapter 4 Laws of Motion 24 Chapter 5 Work Energy and Power 31 Chapter 6 System of Particles and Rotational Motion 39

Chapter 7 Gravitation 47 Chapter 8 Mechanical Properties of Solids 55 Chapter 9 Mechanical Properties of Fluids 62 Chapter 10 Thermal Properties of Matter 69 Chapter 11 Thermodynamics 76 Chapter 12 Kinetic Theory 83 Chapter 13 Oscillations 90 Chapter 14 Waves 98 Chapter 1 Electric Charges and Fields 105 Chapter 2 Electrostatic Potential and Capacitance 113 Chapter 3 Current Electricity 123 Chapter 4 Moving Charges and Magnetism 130 Chapter 5 Magnetism and Matter 138 Chapter 6 Electromagnetic Induction 145 Chapter 7 Alternating Current 153 Chapter 8 Electromagnetic Waves 159 Chapter 9 Ray Optics and Optical Instruments 166 Chapter 10 Wave Optics 174 Chapter 11 Dual Nature of Radiation and Matter 181 Chapter 12 Atoms 188 Chapter 13 Nuclei 194 Chapter 14 Semiconductor Electronics Materials Devices and Simple Circuits 200

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Table of Contents Answers Of Specific Heat

1. Understanding the eBook Answers Of Specific Heat
 - The Rise of Digital Reading Answers Of Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Of Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Of Specific Heat
 - Personalized Recommendations

- Answers Of Specific Heat User Reviews and Ratings
- Answers Of Specific Heat and Bestseller Lists
- 5. Accessing Answers Of Specific Heat Free and Paid eBooks
 - Answers Of Specific Heat Public Domain eBooks
 - Answers Of Specific Heat eBook Subscription Services
 - Answers Of Specific Heat Budget-Friendly Options
- 6. Navigating Answers Of Specific Heat eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Of Specific Heat Compatibility with Devices
 - Answers Of Specific Heat Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Of Specific Heat
 - Highlighting and Note-Taking Answers Of Specific Heat
 - Interactive Elements Answers Of Specific Heat
- 8. Staying Engaged with Answers Of Specific Heat
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of Specific Heat
- 9. Balancing eBooks and Physical Books Answers Of Specific Heat
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of Specific Heat
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Of Specific Heat
 - Setting Reading Goals Answers Of Specific Heat
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Of Specific Heat
 - Fact-Checking eBook Content of Answers Of Specific Heat

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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