

100) Carbon tetrachloride consists of one carbon and four chlorine atoms. How many chlorine atoms are present in 1.00 mole of CCl_4 molecules?

$$\text{one mole of } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl} \text{ atoms}$$

$$1 \text{ mole } \text{CCl}_4 \times \frac{4 \text{ moles Cl}}{1 \text{ mole } \text{CCl}_4} = 4 \text{ moles Cl}$$

$$\text{answer} = 4 \text{ moles Cl} \quad \text{100}$$

101) There are 4 monatomic gases in each buffer gas used for Phosor.

102) Describe how each of these gases is used.

Helium - for cooling the sample
 Argon - for shielding the sample
 Neon - for the sample

These gases are used to cool the sample and to shield it from the environment.

103) Describe how each of these gases is used in the Phosor.

Helium - for cooling the sample
 Argon - for shielding the sample
 Neon - for the sample

104) The Phosor is a compound of carbon and oxygen. The Phosor is a compound of carbon and oxygen.

105) The Phosor is a compound of carbon and oxygen. The Phosor is a compound of carbon and oxygen.

106) The Phosor is a compound of carbon and oxygen. The Phosor is a compound of carbon and oxygen.

$$\text{answer} = 2 \quad \text{106}$$



Answer To F321 Chemistry May Series

DP Hallahan



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