

100 gm solution contains 20 gm H_2SO_4
 $\frac{100}{d}$ ml solution contains 20 gm of H_2SO_4
 or
 1000 ml solution contains 3.6×98 gm of H_2SO_4
 $\therefore 3.6 \times 98 = \frac{20 \times d}{100} = 1000$

$d = 1.23$
 Hence, (3) is correct.

118. The first and second dissociation constants of an acid H_2A are 1.0×10^{-5} and 5.0×10^{-10} respectively. The overall dissociation constant of the acid will be

(1) 5.0×10^{-5} (2) 5.0×10^{-10}
 (3) 5.0×10^{-15} (4) 5.0×10^5

Ans. (3)



For the reaction



$$K = \frac{[H^+]^2[A^{2-}]}{[H_2A]} = K_1 \times K_2$$

$$= 1 \times 10^{-5} \times 5 \times 10^{-10}$$

$$= 5 \times 10^{-15}$$

Hence, (3) is correct.

119. A mixture of ethyl alcohol and propyl alcohol has a vapour pressure of 290 mm at 300 K. The vapour pressure of propyl alcohol is 200 mm. If the mole fraction of ethyl alcohol is 0.6, its vapour pressure (in mm) at the same temperature will be

(1) 350 (2) 300
 (3) 700 (4) 360

Ans. (1)

Sol. Let the vapour pressure of pure ethyl alcohol be P .

According to Raoult's law
 $290 = 200 \times 0.4 + P \times 0.6$

$$P = \frac{290 - 80}{0.6} = 350 \text{ mm Hg}$$

Hence, (1) is correct.

120. In conversion of lime-stone to lime,



the values of ΔH° and ΔS° are $+179.1 \text{ kJ mol}^{-1}$ and 160.2 J/K respectively at 298 K and 1 bar. Assuming that ΔH° do not change with temperature, temperature above which conversion of limestone to lime will be spontaneous is

(1) 1508 K (2) 1200
 (3) 845 K (4) 1158 K

Ans. (4)

Sol. We know, $\Delta G = \Delta H - T\Delta S$

So, let's find the equilibrium temperature, i.e. at which

$$\Delta G = 0$$

$$\Delta H = T\Delta S$$

$$T = \frac{179.1 \times 1000}{160.2}$$

$$= 1158 \text{ K}$$

So, at temperature above this, the reaction will become spontaneous.

Hence, (4) is correct answer.

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Table of Contents Aieee 2paper Solutions Code K

1. Understanding the eBook Aieee 2paper Solutions Code K
 - The Rise of Digital Reading Aieee 2paper Solutions Code K
 - Advantages of eBooks Over Traditional Books
2. Identifying Aieee 2paper Solutions Code K
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aieee 2paper Solutions Code K
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aieee 2paper Solutions Code K
 - Personalized Recommendations
 - Aieee 2paper Solutions Code K User Reviews and Ratings
 - Aieee 2paper Solutions Code K and Bestseller Lists
5. Accessing Aieee 2paper Solutions Code K Free and Paid eBooks
 - Aieee 2paper Solutions Code K Public Domain eBooks
 - Aieee 2paper Solutions Code K eBook Subscription Services
 - Aieee 2paper Solutions Code K Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Aieee 2paper Solutions Code K Compatibility with Devices
 - Aieee 2paper Solutions Code K Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aieee 2paper Solutions Code K
 - Highlighting and Note-Taking Aieee 2paper Solutions Code K
 - Interactive Elements Aieee 2paper Solutions Code K
8. Staying Engaged with Aieee 2paper Solutions Code K
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aieee 2paper Solutions Code K
9. Balancing eBooks and Physical Books Aieee 2paper Solutions Code K
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aieee 2paper Solutions Code K
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Aieee 2paper Solutions Code K
 - Setting Reading Goals Aieee 2paper Solutions Code K
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Aieee 2paper Solutions Code K
 - Fact-Checking eBook Content of Aieee 2paper Solutions Code K
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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