

112. The IUPAC name of



is

- (1) 3, 1-diethyl-4,2,2-dimethylpentane
(2) 4, 4-dimethyl-5, 5-diethylpentane
(3) 5, 5-diethyl-4, 4-dimethylpentane
(4) 3-ethyl-4, 4-dimethylheptane

Ans.
Sol.



The correct answer is 3-ethyl-4, 4-dimethylheptane.

Hence, (4) is correct.

113. Which of the following species exhibits the diamagnetic behaviour?

- (1) O_2^+
(2) O_2
(3) O_2^-
(4) NO

Ans.
Sol.

(1)
The correct option is O_2^+

This species has 18 π , which are filled in such a way that all molecular orbitals are fully filled, so diamagnetic.



Hence, (1) is correct.

114. The stability of divalides of Sn, Ge, Si and Pb increases steadily in the sequence

- (1) $\text{GeX}_2 < \text{SnX}_2 < \text{SnX}_4 < \text{PbX}_2$
(2) $\text{SnX}_2 < \text{GeX}_2 < \text{PbX}_2 < \text{SnX}_4$
(3) $\text{SnX}_2 < \text{GeX}_2 < \text{SnX}_4 < \text{PbX}_2$
(4) $\text{PbX}_2 < \text{SnX}_2 < \text{GeX}_2 < \text{SnX}_4$

Ans.
Sol.

(3)
Due to inert pair effect, the stability of +2 oxidation state increases as we move down this group.
 $\therefore \text{SnX}_2 < \text{GeX}_2 < \text{SnX}_4 < \text{PbX}_2$

Hence, (3) is correct.

115. Identify the incorrect statement among the following

- (1) Ozone reacts with SO_2 to give SO_3
(2) Silicon reacts with $\text{NaOH}_{\text{conc}}$ in the presence of air to give Na_2SiO_3 and H_2O
(3) Cl_2 reacts with excess of NH_3 to give N_2 and HCl
(4) Br_2 reacts with hot and strong NaOH solution to give NaBr , NaBrO_3 and H_2O

Ans.
Sol.

(4)
 Br_2 reacts with hot and strong NaOH to give NaBr , NaBrO_2 and H_2O .

Hence, (4) is incorrect statement.

116. The charge/radius ratio of a cation determines its polarizing power. Which one of the following sequences represents the increasing order of the polarizing power of the cationic species, K^+ , Ca^{2+} , Mg^{2+} , Be^{2+} ?

- (1) $\text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}$
(2) $\text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}$
(3) $\text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}$
(4) $\text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+$

Ans.
Sol.

(3)
Higher the charge/radius ratio, more is the polarizing power.



Hence, (3) is correct.

117. The density (in g mL^{-1}) of a 3.65 M sulphuric acid solution that is 20% H_2SO_4 (Molar mass = 98 g mol^{-1}) by mass will be

- (1) 1.64
(2) 1.88
(3) 1.22
(4) 1.45

Ans.
Sol.

(3)
Let the density of solution be d
Molarity of solution given = 3.6
i.e., 1 litre of solution contains 3.6 moles of H_2SO_4
or 1 litre of solution contains 3.6×98 gms of H_2SO_4
Since, the solution is 20% by mass.

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Mamta Batra



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