

AIR POLLUTION STUDY GUIDE (Quiz 1)

QUIZ 1

Lecture 1:

- **Air pollutant**: contaminant present in the atmosphere at concentrations high enough to adversely affect human health, welfare and/or plant life. Natural and anthropogenic sources. Have sinks and sources.
- Spatial scales (depends on atmospheric lifetime):
 - o Outdoor: local, regional, global
 - o Indoor: home, work
- $N_2 = 78.1\%$ $O_2 = 20.9\%$ $Ar = .9\%$
- $Ppmv = (\text{moles compound}) / (10^6 \text{ moles air})$
- $Mg/m^3 = ppm \times MW \times (P/RT)$
- Ex 1) human health: London Killer Smog Dec 11, 1952
- Ex 2) human health: 4 day pollution episode Denora, PA = 20 deaths + 600 illnesses
- Ex 3) property damage: outdoor statue in Germany circa 1700
- Ex 4) crop damage: acid rain or singular emitter
- Anthropogenic air pollution depends on:
 - o Population
 - o Economic activity per person
 - o Pollutant emissions per unit economic activity
- **Global impact**: gases that spread around the globe (CO_2 , CFCs) w/ long lifetime; unreactive
- **Local impact**: gases/contaminants that react more quickly; short lifetime; (particulates)
- **Primary pollutants**: emitted directly into the atmosphere
- **Secondary pollutants**: formed by rxns in atmosphere; in gas phase or particulates (e.g. ozone)

Lecture 2:

- Ways to control ambient concentrations
 - 1) Limit emissions
 - 2) Limit ambient concentrations
 - 3) Market/emissions trading
- **Clean Air Act**: 40 CFR – contains air pollution laws
- **NAAQS**: National Ambient Air Quality Standards
- **NSPS**: New Source Performance Standards – governs new stationary pollutant sources
- **NESHAP**: National Emissions Standards for Hazardous Pollutants or Air Toxics
- **PSD**: protect areas that are cleaner than NAAQS (visibility in parks)
- **Nonattainment areas**: explains what must be done in areas where NAAQS are not met
- **Mobile sources**: vehicle emissions controlled by federal gov
- **Criteria pollutants**: (6) originally regulated by "Air Quality Criteria" CO , Pb , NO_2 , O_3 , SO_2 , PM

Air Pollution Study Guide Answers

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Rosdahl, Eileen Klein, Mary T. Kowalski**

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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation
 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
 $\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
4 Al + 3 O₂ $\rightarrow$ 2 Al₂O₃
b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.93 \text{ mol Al}$
 $\frac{101.964 \text{ g Al}_2\text{O}_3}{302.3 \text{ g/mol}} = 0.337 \text{ mol Al}_2\text{O}_3$
c determine the percent yield
 $\frac{26.982 \text{ g Al}_2\text{O}_3}{101.964 \text{ g Al}_2\text{O}_3} \times 100 = 26.4\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield $\frac{\text{actual yield g}}{\text{theoretical yield g}} \times 100$ label 3 7 3