

8 • Electron Configurations & Periodicity

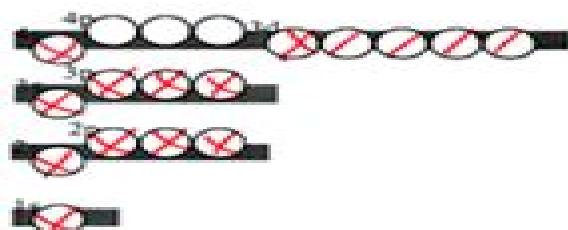
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
 Ar $[\text{Ne}] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
 Mg $[\text{Ne}] 3s^2$
- N $1s^2 2s^2 2p^3$
 N $[\text{He}] 2s^2 2p^3$
- Li $1s^2 2s^1$
 Li $[\text{He}] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
 P $[\text{Ne}] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
 Cl $[\text{Ne}] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$

Fe $[\text{Ar}] 3d^6 4s^2$

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Sebastian Brünink



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