

ABBREVIATIONS AND SYMBOLS			
amount of substance	<i>n</i>	joule	J
atmosphere	atm	kelvin	K
boiling pt. elev. const.	<i>K_b</i>	liter	L
Celsius temperature	°C	molar	<i>m</i>
energy of activation	<i>E_a</i>	molar	M
enthalpy	<i>H</i>	mole	mol
entropy	<i>S</i>	pressure	<i>P</i>
equilibrium constant	<i>K_{eq}</i>	second	s
free energy	<i>G</i>	specific heat capacity	<i>C_p</i>
freezing pt. depr. const.	<i>K_f</i>	temperature	<i>T</i>
gram	g	volt	V
hour	h	volume	V

CONSTANTS	
Planck's constant <i>h</i> = 6.626×10 ⁻³⁴ J·s	Gas Constant <i>R</i> = 0.0821 (L·atm)/(mol·K) <i>R</i> = 8.314 J/(mol·K)
Speed of light <i>c</i> = 2.998×10 ⁸ m·s ⁻¹	Water Dissociation Constant <i>K_w</i> = 1.0×10 ⁻¹⁴ at 25 °C
Rydberg constant <i>R_H</i> = 2.18×10 ⁻¹⁸ J	Faraday Constant <i>F</i> = 96470 C/mol e ⁻ <i>F</i> = 96470 A·s/mol e ⁻ <i>F</i> = 96470 J/V·mol e ⁻
Avogadro Constant <i>N_A</i> = 6.022×10 ²³ mol ⁻¹	
Measures of pressure 1 atm = 760 mmHg	

EQUATIONS			
$E = h\nu$ $E = \frac{hc}{\lambda}$	$\frac{\text{rate}_1}{\text{rate}_2} = \sqrt{\frac{\text{molar mass}_1}{\text{molar mass}_2}}$	Integrated Rate Laws zero: $[A] = [A]_0 - kt$	$\text{pH} = \text{p}K_a + \log\left(\frac{[A^-]}{[HA]}\right)$
$\Delta E = R_{\infty} \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$	$\Delta T_1 = K_f m$ $\Delta T_2 = i K_f m$ $\Delta T_b = K_b m$ $\Delta T_b = i K_b m$ $\Pi = MRT$ $\Pi = i MRT$	second: $\frac{1}{[A]} = \frac{1}{[A]_0} + kt$	$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$
$q = m c \Delta T$ $PV = nRT$	$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right)$	first: $\ln[A] = \ln[A]_0 - kt$	$\Delta G^\circ = -RT \ln K$ $\Delta G = \Delta G^\circ + RT \ln Q$ $E^\circ = \frac{RT}{nF} \ln K$ $E = E^\circ - \frac{RT}{nF} \ln Q$

PERIODIC TABLE OF THE ELEMENTS

PERIODIC TABLE OF THE ELEMENTS																		18	
1 H 1.008													2 He 4.003						
3 Li 6.941	4 Be 9.012													5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 18.99	10 Ne 20.18
11 Na 22.99	12 Mg 24.31													13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.07	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.39	31 Ga 69.72	32 Ge 72.64	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80		
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc 101.1	44 Ru 101.1	45 Rh 102.9	46 Pd 106.4	47 Ag 107.9	48 Cd 112.4	49 In 114.8	50 Sn 118.7	51 Sb 121.8	52 Te 127.6	53 I 126.9	54 Xe 131.3		
55 Cs 132.9	56 Ba 137.3	57 La 138.9	72 Hf 178.5	73 Ta 180.9	74 W 183.8	75 Re 186.2	76 Os 190.2	77 Ir 192.2	78 Pt 195.1	79 Au 197.0	80 Hg 200.6	81 Tl 204.4	82 Pb 207.2	83 Bi 209.0	84 Po 209.0	85 At 210.0	86 Rn 222.0		
87 Fr	88 Ra	89 Ac	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og		
		58 Ce 140.1	59 Pr 140.9	60 Nd 144.2	61 Pm	62 Sm 150.4	63 Eu 152.0	64 Gd 157.3	65 Tb 158.9	66 Dy 162.5	67 Ho 164.9	68 Er 167.3	69 Tm 168.9	70 Yb 173.0	71 Lu 175.0				
		90 Th 232.0	91 Pa 231.0	92 U 238.0	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr				

Answer Key For Acs Divched Examination Institute

K Morrison



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