

Calculating pH

How is pH related to the concentration of hydronium ions?

Why?

In biology and other science courses pH is introduced as a way to quantify the acidity or basicity of a solution. This property can be measured using a pH probe or with an indicator paper strip that changes color at different pH values. But, what is actually being measured? We know that a pH of 7 is neutral, below 7 is acid, and above 7 is base, but why? What in the solution is the paper strip or probe actually reacting with?

Model 1 – Ion Concentrations for Acids and Bases

Beaker	Solution	Acidic, Basic or Neutral?	$[H_3O^+]$	$[OH^-]$	$[H_3O^+] \times [OH^-]$
1	0.10 M HCl(aq)	Acidic	1.0×10^{-1} M	1.0×10^{-13} M	1×10^{-14}
2	0.0010 M HCl(aq)	Acidic	1.0×10^{-3} M	1.0×10^{-11} M	1×10^{-14}
3	0.000010 M HCl(aq)	Acidic	1.0×10^{-5} M	1.0×10^{-9} M	1×10^{-14}
4	0.0000010 M HCl(aq)	Acidic	1.0×10^{-6} M	1.0×10^{-8} M	1×10^{-14}
5	0.00000010 M HCl(aq)	Neutral	1.0×10^{-7} M	1.0×10^{-7} M	1×10^{-14}
6	0.00000010 M NaOH(aq)	Neutral	1.0×10^{-7} M	1.0×10^{-7} M	1×10^{-14}
7	0.0000010 M NaOH(aq)	Basic	1.0×10^{-8} M	1.0×10^{-6} M	1×10^{-14}
8	0.00010 M NaOH(aq)	Basic	1.0×10^{-10} M	1.0×10^{-4} M	1×10^{-14}
9	0.010 M NaOH(aq)	Basic	1.0×10^{-12} M	1.0×10^{-2} M	1×10^{-14}
10	0.10 M NaOH(aq)	Basic	1.0×10^{-13} M	1.0×10^{-1} M	1×10^{-14}

1. What does the symbol $[H_3O^+]$ in Model 1 indicate?

The molar concentration of hydronium ion.

2. In Beaker 2, which ion has a higher concentration, hydronium ion or hydroxide ion?

Hydronium ion.

3. Describe how the concentration of hydronium ion was calculated for Beaker 3 in Model 1 from the concentration of the acid.

Since HCl is a strong acid, all of the molecules ionize in water. Therefore, the H_3O^+ concentration is equal to the initial acid concentration in the solution.

4. Describe how the concentration of hydroxide ion was calculated for Beaker 8 in Model 1 from the concentration of the base.

Since NaOH is a strong base, it completely dissociates in water. Therefore, the OH^- concentration is equal to the initial base concentration in the solution.

Answer Key Pogil Chemistry Calculating Ph

Margaret G. Kimble



Answer Key Pogil Chemistry Calculating Ph:

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Science Citation Index, 1992 Vols for 1964 have guides and journal lists

A Generalised Equation for Calculating PH G. D. Peckham, 1984

Calculation of the PH Value of a Mixture of Solutions Tore K. Gustafsson, 1981

Determination of PH Roger Gordon Bates, 1964 In portraying the rise and fall in eighteenth century Ireland and England of Barry Lyndon an adventurer gambler a cad and a romantic idealist Kubrick departs from Thackeray's picaresque novel in scope and tone The first person narrator of the novel gives way in the film to the third person who assumes a good deal of the storytelling function adding to the sense of detachment and abstraction typical of Kubrick The way that this film polarised the critics suggests that it may hold a key to his oeuvre Enervating pictorialism or a stately meditation upon the trappings of cultural ritual that we call civilisation The painterly tableaux suggest the otherness of a past era a world as alien as that of 2001 in a way matched by few other period films

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