

## Chemical Equations and Reactions

QUIZ  
CH #8  
SECTION #1

## Section Quiz: Describing Chemical Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

B

1. Which of the following is true regarding the law of conservation of mass?
- It dictates that the number of molecules on each side of a chemical equation must be the same.
  - It dictates that the number of atoms of each element must be the same on both sides of a chemical equation.
  - It states that the mass of the reactants must remain constant in order for a chemical reaction to proceed.
  - It does not apply to chemical reactions.

D

2. Which of the following may indicate that a chemical reaction has occurred?
- release of energy as light
  - a color change
  - gas bubble formation
  - All of the above

D

3. Which of the following is a product in the reaction described by the word equation below?



- iron(II) sulfate
- copper
- iron
- Both (a) and (b)

C

4. What does  $\text{CH}_4$  represent in the balanced chemical equation shown below?



- one atom of  $\text{CH}_4$
- five atoms of  $\text{CH}_4$
- one molecule of  $\text{CH}_4$
- 1 g of  $\text{CH}_4$

↑  
1C

5. Which of the following is not true regarding a properly written chemical equation?
- It must be balanced.
  - It must contain proper chemical formulas.
  - It must show the formation of a precipitate or a gas.
  - It should represent known data.

# Answer Key To Section 1describing Chemical Reactions

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