



**Figure 1**

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**Directions (3–10):** For each statement or question, choose the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2014 Edition Reference Tables for Physical Science/Chemistry.

- The oxidation number of nitrogen in  $\text{N}_2$  is  
 (1) +1 (2) +3  
 (3) 0 (4) -3
- What is the oxidation number of Pt in  $\text{K}_2\text{PtCl}_6$ ?  
 (1) -2 (2) -4  
 (3) +2 (4) +4
- In which substance does phosphorus have a +3 oxidation state?  
 (1)  $\text{P}_4\text{O}_{10}$  (2)  $\text{PCl}_5$   
 (3)  $\text{Ca}(\text{PO}_3)_2$  (4)  $\text{KH}_2\text{PO}_4$
- What is the oxidation state of sulfur in  $\text{H}_2\text{SO}_4$ ?  
 (1) 0 (2) -2  
 (3) +6 (4) +4
- In which substance does sulfur have a negative oxidation number?  
 (1)  $\text{Na}_2\text{S}$  (2)  $\text{CaSO}_3$   
 (3)  $\text{S}_8$  (4)  $\text{SO}_2$
- In which compound is the oxidation number of oxygen -1?  
 (1)  $\text{CO}$  (2)  $\text{CO}_2$   
 (3)  $\text{H}_2\text{O}$  (4)  $\text{H}_2\text{O}_2$
- Oxygen has an oxidation number of -2 in  
 (1)  $\text{O}_2$  (2)  $\text{NO}_2$   
 (3)  $\text{Na}_2\text{O}_2$  (4)  $\text{OF}_2$
- Oxygen will have a positive oxidation number when combined with  
 (1) fluorine (2) chlorine  
 (3) bromine (4) iodine
- In which compound does hydrogen have an oxidation number of -1?  
 (1)  $\text{NH}_3$  (2)  $\text{HCl}$   
 (3)  $\text{KH}$  (4)  $\text{H}_2\text{O}$
- Which equation represents an oxidation-reduction reaction?  
 (1)  $\text{HCl} + \text{KOH} \rightarrow \text{KCl} + \text{H}_2\text{O}$   
 (2)  $4\text{HCl} + \text{MnO}_2 \rightarrow \text{MnCl}_2 + \text{H}_2\text{O} + \text{Cl}_2$   
 (3)  $2\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$   
 (4)  $2\text{HCl} + \text{FeS} \rightarrow \text{FeCl}_2 + \text{H}_2\text{S}$
- Which equation represents an oxidation-reduction reaction?  
 (1)  $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$   
 (2)  $\text{Zn}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{ZnCl}_2 + 2\text{H}_2\text{O}$   
 (3)  $\text{H}_2\text{O} + \text{NH}_3 \rightarrow \text{NH}_4^+ + \text{OH}^-$   
 (4)  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{OH}^-$
- Which statement correctly describes a redox reaction?  
 (1) Oxidation and reduction occur simultaneously.  
 (2) Oxidation occurs before reduction.  
 (3) Oxidation occurs after reduction.  
 (4) Oxidation occurs, but reduction does not.

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**Thomas Griffiths**



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