

Molecular Polarity Lab

Name _____ Per. _____

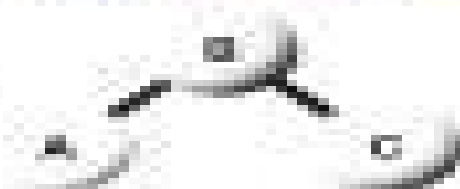
Circle the best correct choice. Check the other correct choice(s) if indicated.

Part A: Three atom investigations

1. Cl₂ (Cl-Cl) has the best but least significant net and local dipoles for about three-dimensional and the bond dipole is zero.
2. Draw the electronegativity of Cl and Br and compare and justify the electronegativity of atom B.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
3. Draw partial dipole(s) for a polar molecule containing the electronegativity of B and answer the results. (P is the following information)
 - a. The polarity arrow points to the _____ electronegativity.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference most _____ the bond dipoles.

Part B: Three atom investigations

1. Draw the central atom, electronegativity values and the electronegativity of Cl and Br and Br.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
2. Draw the partial dipole(s) for the Lewis dot diagram.
 (a) In a different color, describe partial bond dipole direction the diagram.
 (c) Draw the partial charge symbol(s) position on the diagram.
 (d) Circle the electronegative. Add other partial dipole(s) the molecule, for example.
3. Circle the best choice. (a) Draw the "best" arrangement of atoms around the electronegativity of Cl and Br and Br.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
4. Draw the partial dipole(s) for the Lewis dot diagram.
 (a) In a different color, describe partial bond dipole direction the diagram.
 (c) Draw the partial charge symbol(s) position on the diagram.
 (d) Circle the electronegative. Add other partial dipole(s) the molecule, for example.



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DJ Losen



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