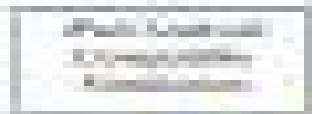


Experiment 1: The Effect of Temperature on the Rate of Reaction

Name: _____

Background Information

Every science experiment is designed to test a hypothesis. In this experiment, the hypothesis is that the rate of reaction will increase as the temperature of the reactants increases. The rate of reaction is the change in concentration of a reactant or product per unit time. The rate of reaction can be measured by the change in concentration of a reactant or product over a given time interval. The rate of reaction can also be measured by the change in volume of a gas produced over a given time interval. The rate of reaction can also be measured by the change in mass of a solid product over a given time interval.



Procedure

$$2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$$



Procedure: 1. Prepare a series of solutions of H_2O_2 at different concentrations.

2. Measure the rate of reaction.

1. Prepare a series of solutions of H_2O_2 at different concentrations. (Using a series of beakers, measure out 100 mL of 1.0 M H_2O_2 and dilute it with water to make 200 mL of 0.5 M H_2O_2 , 300 mL of 0.33 M H_2O_2 , 400 mL of 0.25 M H_2O_2 , 500 mL of 0.2 M H_2O_2 , 600 mL of 0.167 M H_2O_2 , 700 mL of 0.143 M H_2O_2 , 800 mL of 0.125 M H_2O_2 , 900 mL of 0.111 M H_2O_2 , and 1000 mL of 0.1 M H_2O_2 .)
2. Measure the rate of reaction for each solution.
3. Plot the rate of reaction against the concentration of H_2O_2 .

Concentration of H_2O_2 (M)	Volume of O_2 (mL)	Time (s)	Rate of Reaction (mL/s)	Rate of Reaction (mol/L/s)
1.0	100	10	10	0.001
0.5	50	20	2.5	0.00025
0.33	33	30	1.1	0.00011
0.25	25	40	0.625	6.25e-05
0.2	20	50	0.4	4e-05
0.167	16.7	60	0.278	2.78e-05
0.143	14.3	70	0.204	2.04e-05
0.125	12.5	80	0.156	1.56e-05
0.111	11.1	90	0.123	1.23e-05
0.1	10	100	0.1	1e-05

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