

PRACTICE!

LAP SCIENCE



1. Describe one example of how climate affects people.
2. Explain the difference between anecdotal and scientific evidence. Give one example (can be made up) of each for climate change.
3. What does albedo mean?
4. Compare the albedo of an area that is covered in ice to an area that is covered with dark soil.
5. Explain what is meant by the term "Earth's net radiation budget".
6. Distinguish between the angle of inclination and the angle of incidence.
7. Compare and contrast conduction and convection.
8. Describe the Coriolis effect. Which direction does it make currents go in hemisphere?

THERMAL ENERGY TRANSFER IN THE BIOSPHERE

Thermal energy is transferred through the hydrosphere from _____ latitudes to _____ Polar regions largely due to _____ and _____.

SPECIFIC HEAT CAPACITY: the amount of _____ required to raise the temperature of _____ of a substance by _____.

QUANTITY OF THERMAL ENERGY:

formula

$$Q = mc\Delta t$$

EXAMPLES

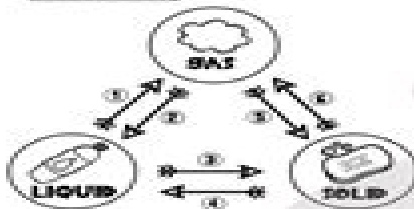
1. A 350-g mass of water at 12.0°C is allowed to warm up to 30.0°C. Determine the amount of thermal energy, Q , absorbed. The theoretical specific heat capacity of water is 4.19 J/g°C.

2. A 100-g mass of iron is at 30°C. Determine the final temperature of the iron after it absorbs 56 J of thermal energy. The theoretical specific heat capacity of iron is 0.449 J/g°C.

usually, when you add energy to a substance, what happens to it or the substance?

but if two specific temperatures compare for both substances, what will that energy go towards instead?

LAP SCIENCE



- 1
- 2
- 3
- 4
- 5
- 6

during phase change, the temperature remains constant

	HEAT OF FUSION	HEAT OF VAPORIZATION
The amount of energy released or absorbed when a substance changes between a...		
Formula		

$$H_{\text{fusion}} = Q = m \cdot n$$

$$H_{\text{vap}} = Q = m \cdot n$$

reminder: moles

formula

$$n = \frac{m}{M}$$

THERMAL ENERGY TRANSFER IN THE BIOSPHERE

specific heat capacity: the amount of _____ required to raise the temperature of _____ of a substance by _____.

specific heat capacity of thermal energy: but if two specific temperatures compare for both substances, what will that energy go towards instead?

$$Q = mc\Delta t$$

Q = thermal energy (J)
 m = mass (g)
 c = specific heat capacity (J/g°C)
 Δt = change in temperature (°C)

$$Q = mc\Delta t$$

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usually, when you add energy to a substance, what happens to it or the substance?

IT gets better

but if two specific temperatures compare for both substances, what will that energy go towards instead?

is phase change

during a phase change, the temperature remains constant?

1: evaporation
2: condensation
3: freezing
4: melting
5: sublimation
6: deposition

	HEAT OF FUSION	HEAT OF VAPORIZATION
The amount of energy released or absorbed when a substance changes between a...		
Formula		

reminder: moles

formula

$$n = \frac{m}{M}$$

Answer Key For Science 10 Alberta

Michael Rajnik

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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web elimu ya sekondari mkoa unazo shule za sekondari 288 ambapo shule za serikali ni 220 na shule binafsi ni 68 aidha shule hizo zote zina jumla ya wanafunzi wa kidato cha i hadi iv 109 432 na wanafunzi wa kidato cha v na vi ni 12 143 kama ilivyo katika jedwali na 53