

MIXED GAS LAWS WORKSHEET - SOLUTIONS

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of $25^\circ C$, what will the pressure of the resulting mixture of gases be? $25^\circ C = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is $20.0^\circ C$. By noon the temperature has risen to $45.0^\circ C$. What is the new volume of the balloon?

$$T_1 = 20.0^\circ C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ C = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in $^\circ C$ of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ C$$

6) $CaCO_3$ decomposes at $1200^\circ C$ to form CO_2 gas and CaO . If 25 L of CO_2 are collected at $1200^\circ C$, what will the volume of this gas be after it cools to $25^\circ C$?

$$T_1 = 1200^\circ C = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ C = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at $20.0^\circ C$ is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is $-20.0^\circ C$?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ C = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ C = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

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papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade. Even though they involve different topics and different systems, they have something in common which can be considered as the signature of the present book. First, these papers are concerned with difficult or very nonideal systems, i.e. systems with very strong interactions, e.g. hydrogen bonding between components or systems with large differences in the partial molar volumes of the components, e.g. the aqueous solutions of proteins or systems that are far from normal conditions, e.g. critical or near critical mixtures. Second, the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures. Last but not least, these systems are of interest for the pharmaceutical, biomedical, and related industries. In order to meet the thermodynamic challenges involved in these complex mixtures, we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and *ab initio* quantum mechanical techniques. The Kirkwood-Buff (KB) theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems. This theory appears to be very fruitful when applied to the above-mentioned difficult systems. The Undersea Journal, 1974 **English Mechanic and World of Science**, 1874

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