

# Using Hess's law to determine enthalpy changes from enthalpy changes of formation.

$$\Delta H_{\text{reaction}} = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$



**Example 6.** What is the enthalpy change for this reaction ?



$$\Delta H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta H = 3 \times \Delta_f H (\text{MgO}) - \Delta_f H (\text{Al}_2\text{O}_3)$$

$$\begin{aligned} \Delta H &= (3 \times -601.7) - -1675.7 \\ &= -129.4 \text{ kJ mol}^{-1} \end{aligned}$$

Remember  
elements have  
 $\Delta_f H = 0$

$$\begin{aligned} \Delta_f H (\text{MgO}) &= -601.7 \text{ kJ mol}^{-1} \\ \Delta_f H (\text{Al}_2\text{O}_3) &= -1675.7 \text{ kJ mol}^{-1} \end{aligned}$$



**Example 7.** Using the following data to calculate the heat of combustion of propene

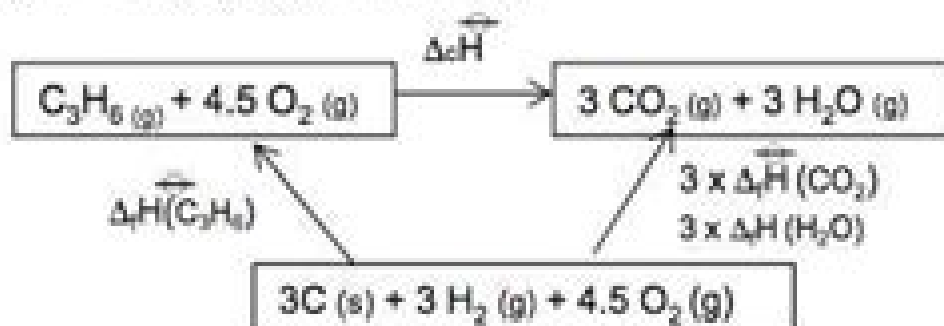
$$\Delta_f H, \text{C}_3\text{H}_6(\text{g}) = +20 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{CO}_2(\text{g}) = -394 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{H}_2\text{O}(\text{g}) = -242 \text{ kJ mol}^{-1}$$



$$\Delta_c H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta_c H = [3 \times \Delta_f H (\text{CO}_2) + 3 \times \Delta_f H (\text{H}_2\text{O})] - \Delta_f H (\text{C}_3\text{H}_6)$$

$$\begin{aligned} \Delta_c H &= [(3 \times -394) + (3 \times -242)] - 20 \\ &= -1928 \text{ kJ mol}^{-1} \end{aligned}$$



# Answers For Chemfax Thermodynamics Enthalpy Lab

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## **Answers For Chemfax Thermodynamics Enthalpy Lab:**

**Paperbound Books in Print 1995** Reed Reference Publishing, Reference Reed, 1995-12      **Commonly Asked Questions in Thermodynamics** Marc J. Assael, Geoffrey C. Maitland, Thomas Maskow, Urs von Stockar, William A. Wakeham, Stefan Will, 2022-08-05 CRC Press is pleased to introduce the new edition of Commonly Asked Questions in Thermodynamics an indispensable resource for those in modern science and engineering disciplines from molecular science engineering and biotechnology to astrophysics Fully updated throughout this edition features two new chapters focused on energy utilization and biological systems This edition begins by setting out the fundamentals of thermodynamics including its basic laws and overarching principles It provides explanations of those principles in an organized manner using questions that arise frequently from undergraduates in the classroom as the stimulus These early chapters explore the language of thermodynamics the first and second laws statistical mechanical theory measurement of thermodynamic quantities and their relationships phase behavior in single and multicomponent systems electrochemistry and chemical and biochemical reaction equilibria The later chapters explore applications of these fundamentals to a diverse set of subjects including power generation with and without fossil fuels for transport industrial and domestic use heating decarbonization technologies energy storage refrigeration environmental pollution and biotechnology Data sources for the properties needed to complete thermodynamic evaluations of many processes are included The text is designed for readers to dip into to find an answer to a specific question where thermodynamics can provide some if not all of the answers whether in the context of an undergraduate course or not Thus its readership extends beyond conventional technical undergraduates to practicing engineers and also to the interested lay person who seeks to understand the discourse that surrounds the choice of particular technological solutions to current and future energy and material production problems      **Thermochemistry and Thermodynamics** Henry Alistair Skinner, 1975      **Hess's Law** Lifeliqe, 2019 This lesson plan covers two indirect methods of determining enthalpy change for most chemical reactions using Hess's law

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