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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-N
Bond energy in kJ/mol	435	193

Use the bond energies in Table 3.1 and $\Delta H = -91 \text{ kJ/mol}$ to calculate the bond energy of an N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4 \times 435) + 1(193)$$

$$4 \times 435 + 1(193) = 2150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond formed} = -91$$

$$2150 - x = -91$$

$$x = 2150 + 91$$

$$x = 2241 \text{ kJ}$$
- Calculate the energy of per mole bond.

$$6 \times \text{N-H} = 2241$$

$$\text{N-H} = \frac{2241}{6} = 373.5 \text{ kJ/mol}$$

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Roman Wölfel



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Table of Contents 0620 31 M J 11 Marking Scheme

1. Understanding the eBook 0620 31 M J 11 Marking Scheme
 - The Rise of Digital Reading 0620 31 M J 11 Marking Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 31 M J 11 Marking Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 31 M J 11 Marking Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 31 M J 11 Marking Scheme
 - Personalized Recommendations
 - 0620 31 M J 11 Marking Scheme User Reviews and Ratings
 - 0620 31 M J 11 Marking Scheme and Bestseller Lists
5. Accessing 0620 31 M J 11 Marking Scheme Free and Paid eBooks
 - 0620 31 M J 11 Marking Scheme Public Domain eBooks
 - 0620 31 M J 11 Marking Scheme eBook Subscription Services
 - 0620 31 M J 11 Marking Scheme Budget-Friendly Options
6. Navigating 0620 31 M J 11 Marking Scheme eBook Formats

- ePub, PDF, MOBI, and More
- 0620 31 M J 11 Marking Scheme Compatibility with Devices
- 0620 31 M J 11 Marking Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 0620 31 M J 11 Marking Scheme
 - Highlighting and Note-Taking 0620 31 M J 11 Marking Scheme
 - Interactive Elements 0620 31 M J 11 Marking Scheme
- 8. Staying Engaged with 0620 31 M J 11 Marking Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 0620 31 M J 11 Marking Scheme
- 9. Balancing eBooks and Physical Books 0620 31 M J 11 Marking Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 0620 31 M J 11 Marking Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 0620 31 M J 11 Marking Scheme
 - Setting Reading Goals 0620 31 M J 11 Marking Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 0620 31 M J 11 Marking Scheme
 - Fact-Checking eBook Content of 0620 31 M J 11 Marking Scheme
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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