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Fig. 3.1

(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	value	unit
bond energy in kJ/mol	440	kJ/mol

Use the bond energies in Table 3.1 and $\Delta H = -2140 \text{ kJ/mol}$ to calculate the bond energy of one C-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 $(1040) + 3(440)$
 $440 + 3(440)$

3320 kJ
- Calculate the energy released when bonds form in the products.
 $\text{bond broken} - \text{bond formed} = -90$
 $3320 - x = -90$
 $x = 3320 + 90$

3410 kJ
- Calculate the energy of C-H bond.
 $6 \times \text{C-H} = 3410$
 $\text{C-H} = \frac{3410}{6}$

568 kJ/mol

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Ying Liu



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Table of Contents 0620 32 M J 09 Mark Scheme

1. Understanding the eBook 0620 32 M J 09 Mark Scheme
 - The Rise of Digital Reading 0620 32 M J 09 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 32 M J 09 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 32 M J 09 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 32 M J 09 Mark Scheme
 - Personalized Recommendations
 - 0620 32 M J 09 Mark Scheme User Reviews and Ratings
 - 0620 32 M J 09 Mark Scheme and Bestseller Lists
5. Accessing 0620 32 M J 09 Mark Scheme Free and Paid eBooks
 - 0620 32 M J 09 Mark Scheme Public Domain eBooks
 - 0620 32 M J 09 Mark Scheme eBook Subscription Services
 - 0620 32 M J 09 Mark Scheme Budget-Friendly Options

6. Navigating 0620 32 M J 09 Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - 0620 32 M J 09 Mark Scheme Compatibility with Devices
 - 0620 32 M J 09 Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 0620 32 M J 09 Mark Scheme
 - Highlighting and Note-Taking 0620 32 M J 09 Mark Scheme
 - Interactive Elements 0620 32 M J 09 Mark Scheme
8. Staying Engaged with 0620 32 M J 09 Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 0620 32 M J 09 Mark Scheme
9. Balancing eBooks and Physical Books 0620 32 M J 09 Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 0620 32 M J 09 Mark Scheme
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 0620 32 M J 09 Mark Scheme
 - Setting Reading Goals 0620 32 M J 09 Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 0620 32 M J 09 Mark Scheme
 - Fact-Checking eBook Content of 0620 32 M J 09 Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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