

- (d) The equilibrium mixture leaving the reaction chamber contains 15% ammonia. Suggest how the ammonia could be separated from the mixture.

	boiling point/ $^{\circ}\text{C}$
hydrogen	-253
nitrogen	-196
ammonia	-33

[2]

- (e) Ammonia is used to make nitrogen trifluoride, NF_3 . Nitrogen trifluoride is essential to the electronics industry. It is made by the following reaction.



Determine if the above reaction is exothermic or endothermic using the following bond energies and by completing the following table. The first line has been done as an example. Bond energy is the amount of energy, in kJ/mole, needed to break or make one mole of the bond.

bond	bond energy in kJ/mole
$\text{N} - \text{H}$	390
$\text{F} - \text{F}$	155
$\text{N} - \text{F}$	280
$\text{H} - \text{F}$	565

bond	energy change / kJ
$\text{N} - \text{H}$	$(3 \times 390) = 1170$
$\text{F} - \text{F}$	
$\text{N} - \text{F}$	
$\text{H} - \text{F}$	

[4]

[Total: 16]

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Table of Contents 0620 33 M J 2014

1. Understanding the eBook 0620 33 M J 2014
 - The Rise of Digital Reading 0620 33 M J 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 33 M J 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 33 M J 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 33 M J 2014
 - Personalized Recommendations
 - 0620 33 M J 2014 User Reviews and Ratings
 - 0620 33 M J 2014 and Bestseller Lists
5. Accessing 0620 33 M J 2014 Free and Paid eBooks
 - 0620 33 M J 2014 Public Domain eBooks
 - 0620 33 M J 2014 eBook Subscription Services
 - 0620 33 M J 2014 Budget-Friendly Options

-
6. Navigating 0620 33 M J 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - 0620 33 M J 2014 Compatibility with Devices
 - 0620 33 M J 2014 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 0620 33 M J 2014
 - Highlighting and Note-Taking 0620 33 M J 2014
 - Interactive Elements 0620 33 M J 2014
 8. Staying Engaged with 0620 33 M J 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 0620 33 M J 2014
 9. Balancing eBooks and Physical Books 0620 33 M J 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 0620 33 M J 2014
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 0620 33 M J 2014
 - Setting Reading Goals 0620 33 M J 2014
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 0620 33 M J 2014
 - Fact-Checking eBook Content of 0620 33 M J 2014
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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