

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|--|---|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 ²⁴ | |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition | |
| _____ the lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula | |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L | |
| _____ the smallest particle of a compound is a compound _____ molar mass | |
| _____ the number representing 6.02 x 10 ²³ representative particles of a substance _____ Avogadro's number | |
| _____ 1 mole = 6.02 ²³ particles _____ empirical formula | |
| _____ the number representing the number of particles of a substance is called _____ 6.02 x 10 ²³ _____ Avogadro's number | |

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following exercise to demonstrate knowledge of the terms and your understanding of the concepts involved with the section of the textbook corresponding to this section (don't give up on math).

Counting whole units (counting, or mass, and or volume for a single quantity called the _____ mass _____). Representation of representative particles (the large unit of measurement is _____ Avogadro's number _____).

Is the mass of 1 mole of a compound, substance, or mixture the _____ percent composition _____ of the atoms making up the compound. Other two techniques for measuring the mass (the smallest unit _____ molar mass _____) of the compound. These are _____ 6.02 x 10²³ _____ representative particles the mass of one substance.

True or False

_____ 1 mole of any substance is 6.02 x 10²³ particles.

_____ 1 mole of a substance contains 6.02 x 10²³ atoms.

_____ 1 mole of a substance contains 6.02 x 10²³ representative particles. **True**

_____ 1 mole of 100 g is composed of one gram of carbon and four atoms of oxygen.

_____ 1 mole of a substance contains 6.02 x 10²³ particles. **True**

Problems

10 Chemical Quantities Chapter Quiz Answer Key

Russell L. Meade



10 Chemical Quantities Chapter Quiz Answer Key:

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Chemical Surety United States. Department of the Army,1995

Chemistry: The Central Science Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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Haynes,Mary Alice Statkiewicz Sherer,Paula J. Visconti,E. Russell Ritenour,2013-12-27 With this workbook you ll enhance your understanding of the material in Radiation Protection in Medical Radiography 6th Edition Author Mary Alice Statkiewicz Sherer uses the same clear accessible approach as in the textbook taking difficult topics and making them easier for you to learn and apply Matching the chapters in the text this workbook ensures that you understand radiation physics and radiation protection and are ready to apply your knowledge in the practice setting Each chapter covers all material included in the text providing a comprehensive review Each chapter highlights important information with an introductory paragraph and a bulleted summary A variety of question formats including matching short discussion items true false multiple choice and fill in the blank questions Calculation exercises offer practice in using formulas and equations presented in the text All answers available in the back of the book so you can easily check your work MCAT Physics and Math Review 2020-2021 Kaplan Test Prep,2019-07-02 Kaplan s MCAT Physics and Math Review 2020 2021 is updated to reflect the latest most accurate and most testable materials on the MCAT A new layout makes our book even more streamlined and intuitive for easier review You ll get efficient strategies detailed subject review and hundreds of practice questions all authored by the experts behind the MCAT prep course that has helped more people get into medical school than all other major courses combined Efficient Strategies and In Depth Review High Yield badges indicate the most testable content based on AAMC materials Concept summaries that boil down the need to know information in each chapter including any necessary equations to memorize Chapter Profiles indicate the degree to which each chapter is tested and the testmaker content categories to which it aligns Charts graphs diagrams and full color 3 D illustrations from Scientific American help turn even the most complex science into easy to visualize concepts Realistic Practice One year online access to instructional videos practice questions and quizzes Hundreds of practice questions show you how to apply concepts and equations 15 multiple choice Test Your Knowledge questions at the end of each chapter Learning objectives and concept checks ensure you re focusing on the most important information in each chapter Expert Guidance Sidebars illustrate connections between concepts and include references to more information real world tie ins mnemonics and MCAT specific tips Comprehensive subject review written by top rated award winning Kaplan instructors who guide you on where to focus your efforts and how to organize your review All material is vetted by editors with advanced science degrees and by a medical doctor We know the test The Kaplan MCAT team has spent years studying every MCAT related document available and our experts ensure our practice questions and study materials are true to the test

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