

Answer **all** the questions in the spaces provided.

- 1 Ethanedioic acid, $\text{HO}_2\text{CCO}_2\text{H}$, has a relative molecular mass of 90.0.

(a) (i) Explain what is meant by the term *relative molecular mass*.

.....

 [2]

(ii) State the empirical formula of ethanedioic acid.

..... [1]

(iii) Calculate how many atoms of carbon are present in 0.18 g of ethanedioic acid, $\text{HO}_2\text{CCO}_2\text{H}$.
 Show your working.

atoms of carbon present = [3]

- (b) Solid ethanedioic acid reacts with aqueous calcium ions to make a precipitate of calcium ethanedioate, CaC_2O_4 .

CaC_2O_4 breaks down when heated to form calcium oxide, carbon dioxide and carbon monoxide.

(i) Construct an equation to represent the reaction of CaC_2O_4 when heated. Include state symbols.

..... [2]

(ii) Identify the type of reaction which occurs when CaC_2O_4 is heated.

..... [1]

(iii) Identify another compound containing calcium ions which will also produce carbon dioxide and calcium oxide when it is heated.

..... [1]

[Total: 10]

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Table of Contents **620 11 Ms M J 09**

1. Understanding the eBook 620 11 Ms M J 09
 - The Rise of Digital Reading 620 11 Ms M J 09
 - Advantages of eBooks Over Traditional Books
2. Identifying 620 11 Ms M J 09
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 620 11 Ms M J 09
 - User-Friendly Interface
4. Exploring eBook Recommendations from 620 11 Ms M J 09
 - Personalized Recommendations
 - 620 11 Ms M J 09 User Reviews and Ratings
 - 620 11 Ms M J 09 and Bestseller Lists

5. Accessing 620 11 Ms M J 09 Free and Paid eBooks
 - 620 11 Ms M J 09 Public Domain eBooks
 - 620 11 Ms M J 09 eBook Subscription Services
 - 620 11 Ms M J 09 Budget-Friendly Options
6. Navigating 620 11 Ms M J 09 eBook Formats
 - ePub, PDF, MOBI, and More
 - 620 11 Ms M J 09 Compatibility with Devices
 - 620 11 Ms M J 09 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 620 11 Ms M J 09
 - Highlighting and Note-Taking 620 11 Ms M J 09
 - Interactive Elements 620 11 Ms M J 09
8. Staying Engaged with 620 11 Ms M J 09
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 620 11 Ms M J 09
9. Balancing eBooks and Physical Books 620 11 Ms M J 09
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 620 11 Ms M J 09
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 620 11 Ms M J 09
 - Setting Reading Goals 620 11 Ms M J 09
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 620 11 Ms M J 09
 - Fact-Checking eBook Content of 620 11 Ms M J 09
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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