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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-H
Bond energy in kJ/mol	436	435

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an N-O bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(400) + 1(193)$$

$$4(436) + 1(498)$$

$$3150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond form} = -92$$

$$3150 = x - 92$$

$$x = 3242 + 92$$

$$3334 \text{ kJ}$$
- Calculate the energy of per N-O bond.

$$6 \times (\text{N-O}) = 3334$$

$$\text{N-O} = \frac{3334}{6}$$

$$555.6 \text{ kJ/mol}$$

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