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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	N-H	N-O
Bond energy in kJ/mol	435	435

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an N-O bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 $(50 \times 4) + 1(4 \times 4)$
 $400 + 1(16)$

2150 kJ
- Calculate the energy released when bonds form in the products.
 $\text{Bond broken} - \text{Bond formed} = -90$
 $2150 - 90 = 2060$

2140 kJ
- Calculate the energy of an N-O bond.
 $6 \times (\text{N-O}) = 2140$
 $\text{N-O} = \frac{2140}{6}$

357 kJ/mol

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Table of Contents 620 63 M J 2013 Mark Scheme

1. Understanding the eBook 620 63 M J 2013 Mark Scheme
 - The Rise of Digital Reading 620 63 M J 2013 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying 620 63 M J 2013 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 620 63 M J 2013 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from 620 63 M J 2013 Mark Scheme
 - Personalized Recommendations
 - 620 63 M J 2013 Mark Scheme User Reviews and Ratings
 - 620 63 M J 2013 Mark Scheme and Bestseller Lists

5. Accessing 620 63 M J 2013 Mark Scheme Free and Paid eBooks
 - 620 63 M J 2013 Mark Scheme Public Domain eBooks
 - 620 63 M J 2013 Mark Scheme eBook Subscription Services
 - 620 63 M J 2013 Mark Scheme Budget-Friendly Options
6. Navigating 620 63 M J 2013 Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - 620 63 M J 2013 Mark Scheme Compatibility with Devices
 - 620 63 M J 2013 Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 620 63 M J 2013 Mark Scheme
 - Highlighting and Note-Taking 620 63 M J 2013 Mark Scheme
 - Interactive Elements 620 63 M J 2013 Mark Scheme
8. Staying Engaged with 620 63 M J 2013 Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 620 63 M J 2013 Mark Scheme
9. Balancing eBooks and Physical Books 620 63 M J 2013 Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 620 63 M J 2013 Mark Scheme
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 620 63 M J 2013 Mark Scheme
 - Setting Reading Goals 620 63 M J 2013 Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 620 63 M J 2013 Mark Scheme
 - Fact-Checking eBook Content of 620 63 M J 2013 Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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