

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পরীক্ষার কপূরা, এখানে কোর্স পরীক্ষা, আরও কয়েকজন শির্ষস্বামীর কলেজসমূহের পরীক্ষা এবং মনোরম টিউটর পদে
পূর্বায় উক্ত অধ্যায়ভিত্তিক সে-একই হয়েছে। এগুলোর অনুশীলন করতে তুমি এ অধ্যায়টি থেকে সেরেবোনা বুঝাননি।
এদের উক্ত লিপিতে পারবে সহজেই।

প্রশ্ন



পাত্র-A



পাত্র-B

- ক. মোলার ভ্রবণ কী?
খ. "কোনো পানিসূত্র লবণের ভ্রাবতা, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
গ. A-পাত্রে বিদ্যমান এলিডের ভর কত নির্ণয় করো।
ঘ. দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অম্লীয় বা ক্ষারীয় হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL ভ্রবণে কোন ভ্রবের 1 mol প্রবীভূত থাকলে উক্ত ভ্রবকে মোলার ভ্রবণ বলে।

কোনো পানিসূত্র লবণের ভ্রাবতা তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে ভ্রাবতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কোনো পানি বৃহৎ লবণ নিষ্কৃতি হয়। তখন ভ্রাবতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে প্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে ভ্রাবতার হ্রাস ঘটে।

অমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
অতএব, প্রবীভূত H_2SO_4 এর ভর = 12.25 g

উদীপকের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
ভ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
ভ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
প্রবীভূত H_2SO_4 এর ভর, $w = ?$

পাত্র-A: H_2SO_4 এর মোল সংখ্যা = $(0.5 \times \frac{250}{1000}) \text{ mol}$
 $= 0.125 \text{ mol}$
পাত্র-B: Na_2CO_3 এর মোল সংখ্যা = $(0.11 \times \frac{250}{1000}) \text{ mol}$
 $= 0.077 \text{ mol}$
সমস্ত বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
সমীকরণ হতে পাই, 1 mol H_2SO_4 = 1 mol Na_2CO_3
1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 0.077 mol " " " " " " " " " " " "
 $(0.077) \text{ mol}$ " " " " " " " " " " " "
 $= 0.077 \text{ mol}$ " " " " " " " " " " " "

অবশিষ্ট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$
পাত্র-A এর ভ্রবণের সাথে পাত্র-B এর ভ্রবণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত ভ্রবণের প্রকৃতি হবে অম্লীয়।

একমল ঘনত্ব ল্যাবরেটরিতে X-পদ্ধতিতে 10 mL এর ঘনত্বের পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর মতুল তৈরি করা করে।

- ক. লোভ কাকে বলে?
খ. Cl_2 অপোলার— ব্যাখ্যা করো।
গ. উদীপকের এলিডটির মোলার ঘনমাত্রা নির্ণয় করো।
ঘ. উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে লোভ বলে।

সমযোজী বন্ধনে আকর্ষণ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা কম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহূণ ঘাটবে।

উদীপকে বর্ণিত এলিডটি হলো 2% $\frac{W}{V}$ HCl।
2% $\frac{W}{V}$ HCl এর অর্থ হলো—
100 mL ভ্রবণে HCl এর ভর 2g
∴ 1000 mL " " " " " " " " " " " "
 $\left(\frac{2 \times 1000}{100} \right) \text{ g}$
 $= 20 \text{ g}$

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Y Pai



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