

Name _____ Date _____ Class _____

CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

Matter—Properties and Changes

Section 3.1 Properties of Matter

In your textbook, read about physical properties and chemical properties of matter.

Use each of the terms below just once to complete the passage.

chemical density	mass properties	physical substance
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Matter is anything with (1) mass and volume. A (2) substance is a form of matter with a uniform and unchanging composition. Substances have specific, unchanging (3) properties that can be observed. Substances have both physical and chemical properties. (4) Physical properties can be observed without changing a substance's chemical composition. Color, hardness, and (5) density are examples. Other properties cannot be observed without changing the composition of a substance. These are called (6) chemical properties. An example is the tendency of iron to form rust when exposed to air.

Label each property as either physical or chemical.

- physical 7. Chemical formula H_2O
- chemical 8. Irons react vigorously when exposed to moist air
- chemical 9. Remains unchanged when in the presence of nitrogen
- physical 10. Colorless
- physical 11. Solid at normal temperatures and pressures
- chemical 12. Ability to combine with another substance
- physical 13. Melting point
- physical 14. Liquid at normal temperatures and pressures
- physical 15. Boiling point is 100°C
- physical 16. Conducts electricity
- physical 17. Density is $\frac{1}{8} \text{ g/cm}^3$

Name _____ Date _____ Class _____

CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

Section 3.1 continued

In your textbook, read about states of matter.

Label each drawing with one of these words: solid, liquid, gas.



For each statement below, write true or false.

- true 21. All matter that we encounter in everyday life exists as one of these physical forms.
- true 22. A solid has definite shape and volume.
- false 23. A liquid has a definite shape and takes on the volume of its container.
- true 24. A gas has both the shape and the volume of its container.
- false 25. The particles in a gas cannot be compressed into a smaller volume.
- false 26. Liquids tend to contract when heated.
- false 27. The particles in a solid are spaced far apart.
- false 28. The words gas and vapor can be used interchangeably.

Answers Chemistry Matter And Change

Edward P. Ortleb, Richard Cadice



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Chemistry: Core Concepts, 3rd Edition Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package

Physical and Chemical Changes (eBook) Edward P. Ortleb, Richard Cadice, 1993-09-01 This book presents a program of basic studies in physical and chemical changes of matter The definition of matter is presented along with explanations of states and properties of matter Topics include atoms molecules elements compounds mixtures solutions symbols and formulas Each of the twelve teaching units in this book is introduced by a color transparency print books or PowerPoint slide eBooks that emphasizes the basic concept of the unit and presents questions for discussion Reproducible student pages provide reinforcement and follow up activities The teaching guide offers descriptions of the basic concepts to be presented background information suggestions for enrichment activities and a complete answer key

Regents Exams and Answers: Chemistry--Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

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Michael Mosher,Paul Kelter,2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry Chemistry For Middle Class-6 (Book-I) R.P. Rana,2010 The multi coloured edition of the books has been revised and updated in a lucid language Project work at the end of each chapter lays emphasis on the learning by doing methodology Unit Test Paper and Model Test Paper have been included **Answer Book for Chemistry, Sixth Edition** Charles E. Mortimer,1986 **The Chemical News** ,1860

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