

9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- ☒ c. transition metal
- d. nonmetal

5. The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the lesser charge, and the suffix *-ic* to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

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Table of Contents Answers Pearson Chemical Names And Formulas Packet

1. Understanding the eBook Answers Pearson Chemical Names And Formulas Packet
 - The Rise of Digital Reading Answers Pearson Chemical Names And Formulas Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Pearson Chemical Names And Formulas Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Pearson Chemical Names And Formulas Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Pearson Chemical Names And Formulas Packet
 - Personalized Recommendations
 - Answers Pearson Chemical Names And Formulas Packet User Reviews and Ratings
 - Answers Pearson Chemical Names And Formulas Packet and Bestseller Lists

5. Accessing Answers Pearson Chemical Names And Formulas Packet Free and Paid eBooks
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 - Answers Pearson Chemical Names And Formulas Packet eBook Subscription Services
 - Answers Pearson Chemical Names And Formulas Packet Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Answers Pearson Chemical Names And Formulas Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Pearson Chemical Names And Formulas Packet
 - Highlighting and Note-Taking Answers Pearson Chemical Names And Formulas Packet
 - Interactive Elements Answers Pearson Chemical Names And Formulas Packet
8. Staying Engaged with Answers Pearson Chemical Names And Formulas Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Pearson Chemical Names And Formulas Packet
9. Balancing eBooks and Physical Books Answers Pearson Chemical Names And Formulas Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Pearson Chemical Names And Formulas Packet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Pearson Chemical Names And Formulas Packet
 - Setting Reading Goals Answers Pearson Chemical Names And Formulas Packet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Pearson Chemical Names And Formulas Packet
 - Fact-Checking eBook Content of Answers Pearson Chemical Names And Formulas Packet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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