

Questions no. 27 to 33 are short answer type questions.

27. Write one chemical equation each for the chemical reaction in which the following have taken place:

- (i) Change in colour
- (ii) Change in temperature
- (iii) Formation of precipitate

Mention colour change/temperature change (rise/fall)/compound precipitated along with the equation. [3]

Solution:

(i) Change in colour:

The reaction between lead nitrate solution and potassium iodide solution.



In this reaction colour changes from colourless to yellow due to formation of PbI_2

(ii) Change in temperature:

The action of dilute sulphuric acid on calcium.



In this reaction, heat evolves.

(iii) Formation of precipitate:

The action of barium chloride on sodium sulphate.



Here, $\text{BaSO}_4(\text{s})$ is a white precipitate.

28. (i) The pH of a sample of tomato juice is 4.6. How is this juice likely to be in taste? Give reason to justify your answer.

(ii) How do we differentiate between a strong acid and a weak base in terms of ion-formation in aqueous solutions?

(iii) The acid rain can make the survival of aquatic animals difficult. How? [3]

Solution:

(i) Tomato juice with a pH of 4.6 is considered slightly acidic. The taste of tomato juice at this pH level is likely to be tangy or slightly sour. This is because acidic substances tend to impart sourness or tanginess to foods and beverages.

(ii) Strong Acids : Ionize fully in water to give H^+ ions in large numbers.

Weak Base : Ionize partially in water to give OH^- ions in small numbers.

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Table of Contents 14 Gradephysical Science Question Paper

1. Understanding the eBook 14 Gradephysical Science Question Paper
 - The Rise of Digital Reading 14 Gradephysical Science Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying 14 Gradephysical Science Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 14 Gradephysical Science Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from 14 Gradephysical Science Question Paper
 - Personalized Recommendations
 - 14 Gradephysical Science Question Paper User Reviews and Ratings
 - 14 Gradephysical Science Question Paper and Bestseller Lists
5. Accessing 14 Gradephysical Science Question Paper Free and Paid eBooks
 - 14 Gradephysical Science Question Paper Public Domain eBooks
 - 14 Gradephysical Science Question Paper eBook Subscription Services
 - 14 Gradephysical Science Question Paper Budget-Friendly Options

6. Navigating 14 Gradephysical Science Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - 14 Gradephysical Science Question Paper Compatibility with Devices
 - 14 Gradephysical Science Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 14 Gradephysical Science Question Paper
 - Highlighting and Note-Taking 14 Gradephysical Science Question Paper
 - Interactive Elements 14 Gradephysical Science Question Paper
8. Staying Engaged with 14 Gradephysical Science Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 14 Gradephysical Science Question Paper
9. Balancing eBooks and Physical Books 14 Gradephysical Science Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 14 Gradephysical Science Question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 14 Gradephysical Science Question Paper
 - Setting Reading Goals 14 Gradephysical Science Question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 14 Gradephysical Science Question Paper
 - Fact-Checking eBook Content of 14 Gradephysical Science Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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