

REVISION FOR PHYSICS A2

1. Which phenomenon is best explained by the particle nature of light?

- A. The Doppler Effect
- B. Polarization
- C. The photoelectric effect
- D. interference

2. According to the quantum theory of light, the energy of light is carried in discrete units called

- A. alpha particles
- B. photoelectrons
- C. protons
- D. photons

3. If the wave properties of a particle are difficult to observe, it is probably due to the particle's

- A. small size
- B. large mass
- C. low momentum
- D. high charge

Ans: larger mass $\Rightarrow$ more complexity $\Rightarrow$ more difficult

4. If an orbiting electron falls to a lower orbit, the total energy of that atom will

- A. remain the same
- B. increase
- C. decrease

Ans: from high to low $\Rightarrow$ release energy via acceleration as real electromagnetic radiation

5. The energy of a photon varies directly with its

- A. frequency
- B. wavelength
- C. speed
- D. rest mass

Ans: $E = h \cdot \nu$ (ν is frequency)

6. Blue light has a frequency of approximately 6.0×10^{14} Hz. A photon of blue light will have an energy of

- A. 5.0×10^{-7} J
- B. 4.0×10^{-19} J
- C. 1.1×10^{-48} J
- D. 6.0×10^{-34} J

Ans: $E = h \cdot \nu$; $h = 6,626 \cdot 10^{-34}$ (ν is frequency)

7. An electron in a hydrogen atom drops from the $n=3$ energy level to the $n=2$ energy level. The energy of the emitted photon is

- A. 4.91eV
- B. 1.51eV
- C. 1.89eV
- D. 3.40eV

Ans : The energy of the emitted photon $\Rightarrow$ find $\Delta E = -13,6 \cdot (1/2^2 - 1/3^2) = -1,88\text{eV}$

8. When an electron changes from a higher energy to a lower energy state within an atom, a quantum of energy is

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