Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:

(a)
$$S_2O_8^{2-} + \Gamma \rightarrow$$

(b)
$$I_2 + S_2O_3^{2-} \rightarrow$$

2) The rate law for the reaction

$$2H_2(g) + 2NO(g) \longrightarrow N_2(g) + 2H_2O(g)$$

is rate = k[H₂][NO]². Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

$$\begin{array}{c} \textit{Mechanism I} \\ \text{H}_2 + \text{NO} \longrightarrow \text{H}_2\text{O} + \text{N} & (\text{slow}) \\ \text{N} + \text{NO} \longrightarrow \text{N}_2 + \text{O} & (\text{fast}) \\ \text{O} + \text{H}_2 \longrightarrow \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

$$\begin{array}{c} \textit{Mechanism II} \\ \text{H}_2 + 2\text{NO} \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O} & (\text{slow}) \\ \text{N}_2\text{O} + \text{H}_2 \longrightarrow \text{N}_2 + \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

$$\begin{array}{c} \textit{Mechanism III} \\ 2\text{NO} \Longrightarrow \text{N}_2\text{O}_2 & (\text{fast equilibrium}) \\ \text{N}_2\text{O}_2 + \text{H}_2 \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O} & (\text{slow}) \\ \end{array}$$

$$\begin{array}{c} \text{N}_2\text{O}_2 + \text{H}_2 \longrightarrow \text{N}_2\text{O}_2 + \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

3) Consider the reaction

From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:

$$CICO_2CCI_3(g) \longrightarrow 2COCI_2(g)$$

Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of CICO ₂ CCl ₃ (mmHg)	
0	15.76	
181	12.64	
513	8.73	
1164	4.44	

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Royston M. Roberts, John C. Gilbert, Stephen F. Martin

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Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07 Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry. It not only highlights the key fundamental components of the coordination chemistry field it also exemplifies the historical development of concepts in the field. In order to graduate as a chemistry major that fills the requirements of the American Chemical Society a student needs to take a laboratory course in inorganic chemistry. Most professors who teach and inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine the environment molecular biology organic synthesis and inorganic materials.

Chemistry

Lemay, Prentice-Hall Staff, 1996

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Gilbert, Stephen F. Martin, 1994 Chemistry McGraw-Hill Staff, 2001-07 Pre-Lab Exercises for Modern

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