

1. The table below shows the number of valence electrons in atoms of elements **W**, **X**, and **Y**.

Element	W	X	Y
No. of valence electrons	1	5	2

- a) Elements **W** and **Y** cannot combine to form a compound. Explain (1 mark)

They both react by losing electrons

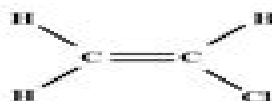
- b) The carbonate of **W** was heated in a test tube, using a Bunsen burner flame. State and explain the observations made. (2 marks)

No production of a colourless gas that forms a white precipitate in calcium hydroxide solution. Carbonates of group 1 elements do not decompose on heating.

2. One mole of ethyne reacted with one mole of hydrogen chloride gas:

- a) Draw the structure and give the name of the compound formed (2 marks)

Structure



Name:

1-chloroethene

- b) Excess hydrogen was bubbled through the product in the presence of nickel catalyst. Three drops of acidified potassium manganate (VII) were then added to the product of reaction. State the observation made when this oxidising agent was added. (1 mark)

Purple acidified potassium manganate (VII) remains purple

3. A certain carbonate has the formula **WCO₃**. If 1g of the carbonate reacts completely with 20cm³ of 1M hydrochloric acid, calculate the relative atomic mass of **W** (3 marks)

(O = 16, C = 12)

Moles HCl	Moles WCO ₃	R.F.M. WCO ₃
$ \frac{20 \times 1}{1000} \\ = 0.02 \text{ moles} $	$ \begin{array}{l} \text{HCl} : \text{WCO}_3 \\ 2 : 1 \\ = \frac{0.02 \times 1}{2} \\ = 0.01 \text{ moles} \end{array} $	$ \begin{array}{l} 0.01 \text{ moles} \equiv 1\text{g} \\ 1 \text{ mole} \equiv \frac{1 \times 1}{0.01} \\ = 100 \\ W + 12 + 3(16) = 100 \end{array} $

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Daniel F McAuley



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