

## PROCEDURE:

Rinse and fill the burette with potassium manganate(VII) solution C-10 ( $\text{KMnO}_4$ ).

Pipette out 20 ml or 25 ml of the oxalic acid solution C-11 ( $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ) in a clean conical flask. To this, add 20 ml of dilute  $\text{H}_2\text{SO}_4$ , C-12, specially provided for this purpose. Warm the contents of the flask to  $60^\circ\text{C}$  –  $70^\circ\text{C}$ . The heating should be continued till the first bubble appears at the bottom of the flask.

Remove the conical flask from fire and titrate this solution by running solution C-10 from the burette. Shake the solution constantly till a permanent pale pink colour is obtained. Ensure that the pink colour obtained does not disappear on shaking the contents of the conical flask.

Repeat the above procedure to get at least two concordant readings.

Tabulate your readings.

State:

- (a) The capacity of the pipette used.
- (b) The titre value you intend to use in your calculations.

Show the titre value to the Visiting Examiner.

The equations for the above reactions are as follows:



Relative atomic masses:



Calculate the following:

- (i) The molarity of oxalic acid solution C-11.
- (ii) The molarity of potassium manganate (VII) solution C-10.
- (iii) The strength of potassium manganate(VII) solution in gms per litre.
- (iv) The percentage purity of the sample of potassium manganate (VII) solution.

*Note: Molarity must be calculated upto at least 4 decimal places.*

# Answer To Alternative B Chemistry Practical 2014 2015

**Jinjun Wang,Huai Liu,Ingeborg  
Menzler-Hokkanen,Hongbo Jiang**



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