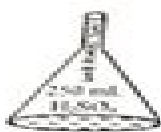


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্লো, এখানে মোট পটাসিয়াম, আরও কপ্লোয়ন শীর্ষস্থানীয় কপ্লোয়নমুকের পটাসিয়াম এবং ময়েন টেম্পের পটাসিয়াম ল্যাবরেটরির অধ্যয়নিক সেএক্স হলেখে। এপুলের অনুবীলন করলে তুমি এ অধ্যয়নিক থেকে সেএক্সের ল্যাবরেটরির অধ্যয়নিক সেএক্স হলেখে।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেন্দ্র পটাসিয়াম ল্যাবরেটরির প্রকণতা, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পটাসিয়াম বৃদ্ধি না পেয়ে প্রকণ পাত্র'— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা প্রকণ হবে— বিয়োজন করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলের প্রকণ বলে।

কেন্দ্র পটাসিয়াম ল্যাবরেটরির প্রকণতা অপমাত্রা বৃদ্ধির সাথে সাথে এক পটাসিয়াম বৃদ্ধি না পেয়ে প্রকণ পাত্র। এর কারণ, অপমাত্রা বৃদ্ধির সাথে সাথে প্রকণতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট অপমাত্রার কেন্দ্র পটাসিয়াম ল্যাবরেটরির প্রকণিত হয়। তখন প্রকণতা প্রকণ পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর প্রকণ বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কেন্দ্র পটাসিয়ামে প্রকণিত করার সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃদ্ধি কিন্তু 32.4°C অপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ প্রকণিত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C অপমাত্রার উপরে অপমাত্রা বৃদ্ধিজনিত কারণে প্রকণতার প্রকণ হয়।

প্রশ্ন ২

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
 অতএব, প্রকণিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
 প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
 প্রকণিত H_2SO_4 এর ভর, $w = ?$

প্রশ্ন-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রশ্ন-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতি বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
 সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

প্রশ্ন-A এর প্রকণের সাথে প্রশ্ন-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকণি হবে অক্সিজেন।

প্রশ্ন ২ এককাল ছাত্র ল্যাবরেটরিতে X-পটাসিয়ামে 10 mg এর প্রকণ প্রকণ এবং Y-পটাসিয়ামে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর প্রকণ প্রকণ করে।

সি. কে. ১০/১০/১০

- নোট কাকে বলে?
- Cl_2 অপোলায়— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পটাসিয়ামে পরিবেশবান্ধব? বিয়োজন করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন। অপরদিকে, আয়নিক বন্ধনে আবদ্ধ পরমাণুর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর বেশি হবে অপরদিকে হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C এর তড়িৎ ঋণাত্মকতা 2.5 তাই $(3 - 2.5) = 0.5$, তাই Cl_2 বন্ধন সমযোজী ও অপরদিকে যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এতদূর থাকে।

উদাহরণে বর্ণিত এলিটটি হলো 2% $\left(\frac{W}{V}\right)$ HCl।

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

Answer Of Hsc Chemistry 1st Paper 15 Ctg Board

2018 Edit, Mark Dixon



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