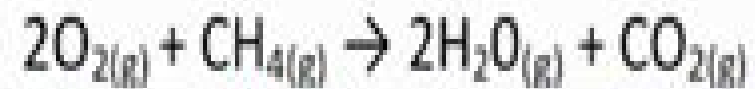


Use Hess' Law(s) for the following reaction to: (a) determine whether exothermic/endothermic, (b) find the change in entropy and (c) determine whether spontaneous or not. Check your answer to (c) using your results from (a) and (b) and Gibbs free energy equation.



$$\Delta H^{\circ}_{\text{reaction}} = \sum \Delta H^{\circ}_f(\text{products}) - \sum \Delta H^{\circ}_f(\text{reactants})$$

$$\Delta H^{\circ}_{rxn} = 2 \cdot (-241,8) - 393,5 - (2 \cdot 0 - 74,8)$$

$$= -802,3 \frac{\text{kJ}}{\text{mol}} < 0 \text{ exothermic}$$

0	-74,8	-241,8	-393,5	$\frac{\text{kJ}}{\text{mol}}$
205,1	186,3	188,8	213,7	$\frac{\text{J}}{\text{mol} \cdot \text{K}}$

$$\Delta S^{\circ}_{\text{reaction}} = \sum S^{\circ}(\text{products}) - \sum S^{\circ}(\text{reactants})$$

$$\Delta S^{\circ}_{rxn} = 2 \cdot 188,8 + 213,7 - (2 \cdot 205,1 + 186,3) = -5,2 \frac{\text{J}}{\text{mol} \cdot \text{K}}$$

$$\Delta G^{\circ}_{\text{reaction}} = \sum \Delta G^{\circ}_f(\text{products}) - \sum \Delta G^{\circ}_f(\text{reactants})$$

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