

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals in the $3p$ sublevel have more energy than orbitals in the $3s$ sublevel. The following orbitals have sublevels from lowest to highest energy:



Finally, you have information for which orbitals are filled. Here's the orbital configuration for orbitals:



Not all orbitals appear in the list of orbitals. This is because the energy of orbitals is not the same. For example, the $3d$ orbitals have higher energy than the $4s$ orbitals, so the $4s$ orbitals are filled first, then the $3d$ orbitals.

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will occupy a $3d$ orbital only after the $4s$ orbitals are filled, and a $4s$ orbital will only occupy after the $3d$ orbitals are filled. Additionally, two electrons can occupy a $3d$ orbital. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy the same orbital, they must spin in opposite directions—orbital spins are opposite directions.

Critical Thinking Questions

- a. How many orbitals would a sublevel have if it had 10 orbitals? (This is called the **lowest energy principle**.)
b. How many orbitals would a sublevel have if it had 10 orbitals? (This is called the **lowest energy principle**.)
c. How many orbitals would a sublevel have if it had 10 orbitals? (This is called the **lowest energy principle**.)
- a. How many orbitals would a sublevel have if it had 10 orbitals? (This is called the **lowest energy principle**.)
b. How many orbitals would a sublevel have if it had 10 orbitals? (This is called the **lowest energy principle**.)
c. How many orbitals would a sublevel have if it had 10 orbitals? (This is called the **lowest energy principle**.)

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin.



Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin.



Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin. Electrons in the same orbital will have the same spin.

The following are the orbital configurations for the elements, nitrogen and oxygen. Write the orbital configuration for each element in the space provided. The $3d$ orbitals are filled first, then the $4s$ orbitals are filled.



Critical Thinking Questions

- a. How many "upspin" electrons are in a sublevel? (This is called the **lowest energy principle**.)
b. How many "upspin" electrons are in a sublevel? (This is called the **lowest energy principle**.)
c. How many "upspin" electrons are in a sublevel? (This is called the **lowest energy principle**.)
- a. How many "upspin" electrons are in a sublevel? (This is called the **lowest energy principle**.)
b. How many "upspin" electrons are in a sublevel? (This is called the **lowest energy principle**.)
c. How many "upspin" electrons are in a sublevel? (This is called the **lowest energy principle**.)

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**Corwin Hansch, Peter George
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