

CHAPTER 10: Solutions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The standard heat decomposition of a substance is determined by the following thermodynamic equation:



- a. Describe how the value of $\Delta H_{\text{decomp}}^\circ$ is being interpreted (sign of $\Delta H_{\text{decomp}}^\circ$).
- b. What is the correct sign for the standard change in the decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

3. _____ $\Delta H_{\text{decomp}}^\circ$ of a substance is 20.4° kJ/mol. If a sample of 4.00 g of substance is heated until it decomposes, how much heat is absorbed?

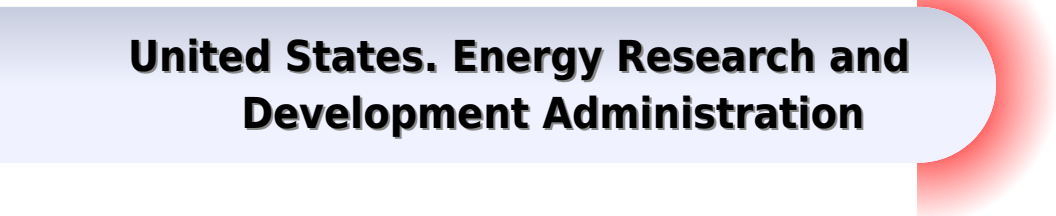
4. _____ A substance has a specific heat of 0.001 J/g $^\circ\text{C}$. If a sample weighing 10.0 g is heated from 25°C to 100°C , how much heat is absorbed?

5. The product is a molecular compound of 10.0 g/mol, and the standard heat of formation of this compound is _____.

- a. What is the value of ΔH_f° for this reaction?

Answers For Reaction Energy Section 1

**United States. Energy Research and
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