

5 (d) Explain why the current in the external circuit of this cell falls to zero after the cell has operated for some time.

.....

.....

(1 mark)

5 (e) The simplified electrode reactions in a rechargeable lithium cell are

Electrode A $\text{Li}^+ + \text{MnO}_2 + \text{e}^- \longrightarrow \text{LiMnO}_2$ $E = -0.15 \text{ V}$

Electrode B $\text{Li}^+ + \text{e}^- \longrightarrow \text{Li}$

Electrode B is the negative electrode.

5 (e) (i) The e.m.f. of this cell is 2.90 V.

Use this information to calculate a value for the electrode potential of electrode B.

.....

.....

(1 mark)

5 (e) (ii) Write an equation for the overall reaction that occurs when this lithium cell is being recharged.

.....

.....

.....

(2 marks)

5 (e) (iii) Suggest why the recharging of a lithium cell may lead to release of carbon dioxide into the atmosphere.

.....

.....

(1 mark)

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Table of Contents Answers For Stars Suite Chemistry

1. Understanding the eBook Answers For Stars Suite Chemistry
 - The Rise of Digital Reading Answers For Stars Suite Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Stars Suite Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Stars Suite Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Stars Suite Chemistry

- Personalized Recommendations
- Answers For Stars Suite Chemistry User Reviews and Ratings
- Answers For Stars Suite Chemistry and Bestseller Lists
- 5. Accessing Answers For Stars Suite Chemistry Free and Paid eBooks
 - Answers For Stars Suite Chemistry Public Domain eBooks
 - Answers For Stars Suite Chemistry eBook Subscription Services
 - Answers For Stars Suite Chemistry Budget-Friendly Options
- 6. Navigating Answers For Stars Suite Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Stars Suite Chemistry Compatibility with Devices
 - Answers For Stars Suite Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Stars Suite Chemistry
 - Highlighting and Note-Taking Answers For Stars Suite Chemistry
 - Interactive Elements Answers For Stars Suite Chemistry
- 8. Staying Engaged with Answers For Stars Suite Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Stars Suite Chemistry
- 9. Balancing eBooks and Physical Books Answers For Stars Suite Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Stars Suite Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Stars Suite Chemistry
 - Setting Reading Goals Answers For Stars Suite Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Stars Suite Chemistry

- Fact-Checking eBook Content of Answers For Stars Suite Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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