CHEM 151 – Quiz #4 Thermo, Redox, and Electrochem December 4, 2018

1. Consider this unbalanced reaction at 25 °C:

Zinc solid reacts with a solution containing nitrate ion, to form zinc ion and nitrogen monoxide gas.

The standard reduction potentials (E°) are:

a. (2 pts) Using chemical formulas, write the unbalanced reaction given above.

 (5 pts) Balance the reaction in an <u>acidic</u> solution using the half-reaction method practiced in class (include phases in your final balanced reaction).

$$3(2n \rightarrow 2n^{2+} + 2e^{-})$$
 $2(3e^{-}+4H^{+} + N05^{-} \rightarrow N0 + 2H_{2}0)$
 $32n + 9e^{-} + 8H^{+} + 2N05^{-} \rightarrow 32n^{2+} + 6e^{-} + 2N0 + 4H_{2}C$

c. (3 pts) i) assign oxidation numbers to all atoms in your final balanced reaction, ii) calculate the change of oxidation state for the two atoms that underwent change, and identify if each atom underwent oxidation or reduction, and iii) label the oxidizing and reducing agents (from the reactants only).

ii)
$$Zn \not \otimes \longrightarrow +z$$
, oxidation $\Longrightarrow Zn \otimes = reducing agent$
 $N +5 \longrightarrow +z$, reduction $\Longrightarrow NO_{\overline{S}(ag)} = oxidizing agent$

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Ian Pickup

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