

Equilibrium Simulation Lab

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Date: 3-2-2013

1. Temperature and Molecular Motion: Decrease the temperature to 100K. Observe the motion of 2-3 gas molecules. Repeat then increase the temperature to 300K. Again observe the motion. Enable tracking, and observe the highlighted molecules.

a. What is the difference in the speed of the molecules at the different temperatures?

- The higher the temp, the faster the molecules move.

b. At a given temperature, does the tracked molecule always move at the same speed?

- The speed of the tracked molecule depends on the speed of the other molecules that interact with it.

2 Reaction: Enable tracking, select oxygen on the top blue menu, and enable tracking. Allow the reaction to run for a few seconds, until the mixture changes. Freeze, open the window, click back to the tracking, and step through until you only see the molecules react. Freeze, enable reactions, and restart. Observe the colored, bonded molecules.

a. What is the reaction that occurred? Give the balanced reaction from memory, noting the relative subscripts, i.e., $H_2 + O_2 \rightarrow H_2O$. Note that the equation adds the fact that not only the all H₂ and O₂ molecules.



b. Do all collisions between H₂ and O₂ molecules cause a reaction? Why or why not?

- No. When an H₂ molecule collides with O₂ molecule or 2x O₂ will, a reaction takes place. The orientation of the O₂ molecule determines if a reaction takes place.

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