

100. Determine the number of molecules of SO_2 molecules. Show SO_2 molecule is a trigonal planar. How many atoms of sulfur are there in 1.00 mole of SO_2 molecules?

$$\text{no. of SO}_2 \text{ molecules} = 1.00 \text{ mol} \times (6.02 \times 10^{23}) \text{ molecules/mol} = 6.02 \times 10^{23} \text{ molecules}$$

$$\text{atoms of } \text{SO}_2 = 3 \times 6.02 \times 10^{23} = 1.81 \times 10^{24} \text{ atoms} \quad 100$$

101. The only commercially known diatomic sulfur are van der Waals.

101. Describe how van der Waals forces arise.

Electron clouds are attracted to each other. The electron clouds are attracted to each other. The electron clouds are attracted to each other.

London dispersion forces are a type of van der Waals force. They arise from the temporary dipoles that are induced in atoms and molecules.

102. Suggest why there are no other intermolecular forces in solid sulfur.

There are no other intermolecular forces in solid sulfur because sulfur is a non-polar molecule and only London dispersion forces are present.

103. Sulfur dioxide is a component of sulfur dioxide. Suggest why sulfur dioxide is a gas at room temperature.

Sulfur dioxide is a gas at room temperature because the forces between SO_2 molecules are weak.

What is the relative number of sulfur in SO_2 , SO_3 , SO_4^{2-} ?



Answer To F321 Chemistry May Series

Yin Toon Tan



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