UNIVERSITY OF BOTSWANA

Deteriol: 3 Selections

 Explain why for a given alkali metal the second ionization energy is much higher than the first ionization energy.

Serbattion.

Removal of the first electron (first tomastion) is relatively easy as the electron is taken from the outer shell build-filled orbital. Removal of the second electron is from a filled atomic orbital, and this involves a great deal of energy because the electron is being removed from a noble gas-configuration. For Li, this is from 1s² orbital; for beavier group 1 metals, it is from an (n-1)p² level. Furthermore, the nucleus attraction is high in a carion than in a neutral atom.

All the group 1 metals react violently with water and the intensity of the reaction increases as the group is descended. Explain.

Section from

Given that group I metals have I electron in the outer shell, they will react easily by donating that electron. This ability to donate the single electron increases as the electron becomes bigger because the electron is situated for from the nucleus bence lower ionization energy and high reactivity for large group I elements.

Suggest what will happen when a mixture of Lift and Naff is heated. The following information is provided: AdV(Lift, S) = -270 KJ mo?* AdV(Naft, S) = -288 KJ mo?* AdV(Lift, S) = -616 KJ mo?* AdV(Naft, S) = -577 KJ mo?*.

Solution.

The possible resented that could occur in:

40140E-221

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