

PROCEDURE:

Rinse and fill the burette with potassium manganate(VII) solution C-10 (KMnO_4).

Pipette out 20 ml or 25 ml of the oxalic acid solution C-11 ($\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$) in a clean conical flask. To this, add 20 ml of dilute H_2SO_4 , C-12, specially provided for this purpose. Warm the contents of the flask to 60°C – 70°C . The heating should be continued till the first bubble appears at the bottom of the flask.

Remove the conical flask from fire and titrate this solution by running solution C-10 from the burette. Shake the solution constantly till a permanent pale pink colour is obtained. Ensure that the pink colour obtained does not disappear on shaking the contents of the conical flask.

Repeat the above procedure to get at least two concordant readings.

Tabulate your readings.

State:

- (a) The capacity of the pipette used.
- (b) The titre value you intend to use in your calculations.

Show the titre value to the Visiting Examiner.

The equations for the above reactions are as follows:



Relative atomic masses:



Calculate the following:

- (i) The molarity of oxalic acid solution C-11.
- (ii) The molarity of potassium manganate (VII) solution C-10.
- (iii) The strength of potassium manganate(VII) solution in gms per litre.
- (iv) The percentage purity of the sample of potassium manganate (VII) solution.

Note: Molarity must be calculated upto at least 4 decimal places.

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M Carnoy



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1. Understanding the eBook Answer To Alternative B Chemistry Practical 2014 2015
 - The Rise of Digital Reading Answer To Alternative B Chemistry Practical 2014 2015
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer To Alternative B Chemistry Practical 2014 2015
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer To Alternative B Chemistry Practical 2014 2015
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer To Alternative B Chemistry Practical 2014 2015
 - Personalized Recommendations
 - Answer To Alternative B Chemistry Practical 2014 2015 User Reviews and Ratings
 - Answer To Alternative B Chemistry Practical 2014 2015 and Bestseller Lists
5. Accessing Answer To Alternative B Chemistry Practical 2014 2015 Free and Paid eBooks
 - Answer To Alternative B Chemistry Practical 2014 2015 Public Domain eBooks
 - Answer To Alternative B Chemistry Practical 2014 2015 eBook Subscription Services
 - Answer To Alternative B Chemistry Practical 2014 2015 Budget-Friendly Options
6. Navigating Answer To Alternative B Chemistry Practical 2014 2015 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer To Alternative B Chemistry Practical 2014 2015 Compatibility with Devices
 - Answer To Alternative B Chemistry Practical 2014 2015 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer To Alternative B Chemistry Practical 2014 2015
 - Highlighting and Note-Taking Answer To Alternative B Chemistry Practical 2014 2015
 - Interactive Elements Answer To Alternative B Chemistry Practical 2014 2015
8. Staying Engaged with Answer To Alternative B Chemistry Practical 2014 2015

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer To Alternative B Chemistry Practical 2014 2015
- 9. Balancing eBooks and Physical Books Answer To Alternative B Chemistry Practical 2014 2015
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer To Alternative B Chemistry Practical 2014 2015
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer To Alternative B Chemistry Practical 2014 2015
 - Setting Reading Goals Answer To Alternative B Chemistry Practical 2014 2015
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer To Alternative B Chemistry Practical 2014 2015
 - Fact-Checking eBook Content of Answer To Alternative B Chemistry Practical 2014 2015
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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