

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals in the $3p$ sublevel have more energy than orbitals in the $3s$ sublevel. The following orbitals have different energy levels in highest energy.



Circle the orbitals that have the same energy level. (Remember, orbitals in the same energy level are called orbitals.)



Now list the orbitals that have the same energy level. (Remember, orbitals in the same energy level are called orbitals.)

When orbitals have the same energy level, they are called orbitals of the same energy level. (This is called the **Pauli Exclusion Principle**.) An orbital can only hold two electrons with opposite spins. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy the same orbital, they must spin in opposite directions—orbital spins must be opposite.

Critical Thinking Questions

- a. How many orbitals would a sublevel have if it had 10 orbitals? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)
b. How many orbitals would a sublevel have if it had 10 orbitals? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)
- a. How many orbitals would a sublevel have if it had 10 orbitals? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)
b. How many orbitals would a sublevel have if it had 10 orbitals? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". (Remember, orbitals in the same energy level are called orbitals of the same energy level.)



Remember, orbitals in the same energy level are called orbitals of the same energy level. (Remember, orbitals in the same energy level are called orbitals of the same energy level.)



Remember, orbitals in the same energy level are called orbitals of the same energy level. (Remember, orbitals in the same energy level are called orbitals of the same energy level.)

The following are the orbitals in the same energy level. (Remember, orbitals in the same energy level are called orbitals of the same energy level.)



Critical Thinking Questions

- a. How many "upspin" orbitals are in a sublevel? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)
- a. How many "upspin" orbitals are in a sublevel? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)
- a. How many "upspin" orbitals are in a sublevel? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)
- a. How many "upspin" orbitals are in a sublevel? (Remember, orbitals in the same energy level are called orbitals of the same energy level.)

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Christian Drosten



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