AP Chemistry Chapter 4 Practice Exam

- The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

$$2H_2S + 3O_2 \rightarrow 2SO_2 + 2H_2O$$

 $SO_2 + 2H_2S \rightarrow 3S + 2H_2O$

How much sulfur (in grams) is produced from 21.5 grams of O2? 21.5502 | mol 02 | 2502 | 35 | 32.0675 = 43.155

- A) 97.0 g
- B) 14.4 g
- C) 21.5 gD) 43.1 g
- none of these
- B
- Iron is produced from its ore by the reactions:

$$2C(g) + O_2(g) \rightarrow 2CO(g)$$

$$\text{Fe}_2\text{O}_3(\varepsilon) + 3\text{CO}(g) \rightarrow 2\text{Fe}(\varepsilon) + 3\text{CO}_2(g)$$

How many moles of $O_2(g)$ are needed to produce 4.6 moles of Fe(s)?

- A) 2.3 mole O₂ B) 3.5 mole O₂ C) 4.6 mole O₂
 - D) 6.9 mole O₂
 - E) 13.8 mole O₂

- B
- Suppose the reaction Ca₂(PO₄)₂ + 3H₂SO₄ → 3CaSO₄ + 2H₂PO₄ is carried out starting with 153 g of Ca₁(PO₄)₂

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 gB) 51.1 g $153 g Ca_3 Pdq)_2 | 1m_1 (s_3 Cu)_2 | 2 H_3 Pdq$ $1 Ca_3 Cu)_3 | 1m_1 H_3 Cu_4$

- (B) 51.1 g
- C) 229.8 g D) 115.1 g
- E) 96.7 g
- 7 6.89 H2 SOY | Ind H, SOY | 2 H2 POY 1 2 =

- SO₂ reacts with H₂S as follows:

$$2H_2S + SO_2 \rightarrow 3S + 2H_2O$$

When 7.50 g of H₂S reacts with 12.75 g of SO₂, which statement applies?

- A) 6.38 g of sulfur are formed.
- B) 10.6 g of sulfur are formed.
 - C) 0.0216 moles of H₂S remain.
 - D) 1.13 g of H₂S remain.
- E) SO₂ is the limiting reagent.

4.6 mil Fersi 300 102 = 3,45m0102

Answers For Ap Chemistry Test Open Ended

CL Gary

Answers For Ap Chemistry Test Open Ended:

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