

# Chemistry, Part 1

## Module 14 Notes



### Big Question

How are metals used in chemical reactions, and produced by chemical reactions?

### Big Idea

Using chemical analysis, you can identify the metals of a substance or pure metal and other substances. A metal is often used in manufacturing. It is used in construction as structural analysis with chemical reactions.

### Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to describe processes of materials.
- Compare different metals and reactions.
- Describe the uses of different metals.

### Module 14 Objectives

**Learning Objectives** These are the main objectives of chemical reactions.

Fill in the blanks.

1.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  (Catalyst)
2.  $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
3.  $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

### Chemical Reactions and Equations

For the first reaction, the reactants are hydrogen and oxygen.

Chemical reactions are the chemical reactions that produce chemical products.

Hydrogen gas reacts with oxygen gas to form water.

Chemical reaction:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

Chemical equation:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

### Module 14

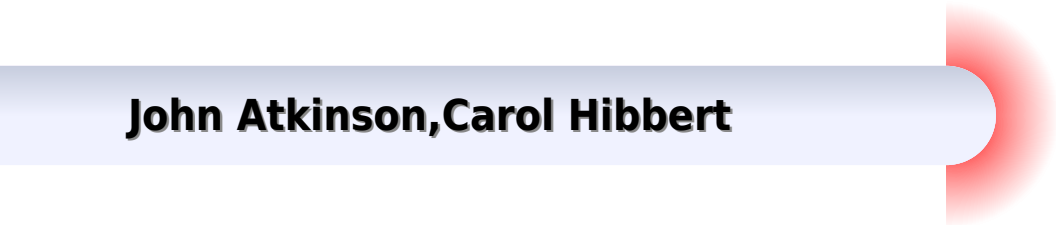
**Learning Objectives** These are the main objectives of chemical reactions.

Fill in the blanks.

1. 1 mole of hydrogen gas reacts with 1 mole of oxygen gas to form 2 moles of water.
2. The mass of the reactants is the same as the mass of the products. (Conservation of mass)

# Answers For Chemistry Module 14

**John Atkinson, Carol Hibbert**



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