

Types of Chemical Reaction Worksheet

Balance the reactions **1 to 6** and indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented:

- $2 \text{NaBr} + \text{Ca(OH)}_2 \rightarrow \text{CaBr}_2 + 2 \text{NaOH}$ Reaction Type : single displacement
- $2 \text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow 3 (\text{NH}_4)_2\text{SO}_4$ Reaction Type : synthesis
- $4 \text{C}_3\text{H}_8\text{O} + 27 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$ Reaction Type : combustion
- $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$ Reaction Type : single displacement
- $\text{Li}_3\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$ Reaction Type : double displacement
- $\text{HBr} + 3 \text{Al(OH)}_3 \rightarrow \text{H}_2\text{O} + 3 \text{AlBr}_3$ Reaction Type : double displacement

Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented in 7 to 20.

- $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ Reaction Type **double displacement**
- $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ Reaction Type **double displacement**
- $\text{C}_8\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$ Reaction Type **combustion**
- $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ Reaction Type **single displacement**
- $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ Reaction Type **decomposition**
- $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$ Reaction Type **synthesis**
- $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + 2 \text{RbF}$ Reaction Type **double displacement**
- $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + 2 \text{Ag}$ Reaction Type **single displacement**
- $\text{C}_3\text{H}_8\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ Reaction Type **combustion**
- $2 \text{C}_3\text{H}_6 + \text{Fe} \rightarrow \text{Fe(C}_3\text{H}_5)_2$ Reaction Type **synthesis**
- $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ Reaction Type **single displacement**
- $2 \text{MgI}_2 + \text{Mn(SO}_4)_2 \rightarrow 2 \text{MgSO}_4 + \text{MnI}_4$ Reaction Type **double displacement**
- $\text{O}_3 \rightarrow \text{O} + \text{O}_2$ Reaction Type **decomposition**
- $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$ Reaction Type **decomposition**

Answers For Chemical Reactions

V. K. Sally



Answers For Chemical Reactions:

Unveiling the Magic of Words: A Report on "**Answers For Chemical Reactions**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Answers For Chemical Reactions**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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