

Atomic Structure

Atoms are made from protons, neutrons, and electrons. Protons and neutrons are known as nucleons.

Electrons have a negligible mass compared to protons and neutrons which both have a relative mass of 1. Protons and electrons have equal and opposite magnitudes of charge, $1+$, $1-$.

MASS NUMBER – Total number of protons and neutrons in the nucleus

ATOMIC NUMBER – The number of protons in the nucleus of an atom.

Ions are formed when atoms lose or gain electrons.

ISOTOPE – An atom of the same element with a different number of neutrons.

Isotopes have the same chemical properties as they have the same electron configuration as one another. They have different physical properties to one another as this is due to the variance in mass.

The model of the atom has changed over time due to the increase in understanding:

- Dalton's model: Suggests that atoms are solid spheres and different spheres make up different elements
- Thomson's Model: Atoms are a positive charge with negatively charged particles in them.
- Rutherford's Model: Small positive nucleus in the centre of the nucleus and electrons surrounding it, mostly empty space
- Bohr's Model: Electrons exist in fixed orbitals and these have fixed energy.

RELATIVE ATOMIC MASS- Average mass of an atom of an element on a scale where $1/12^{\text{th}}$ of C^{12} is exactly 1.

Steps of mass spectrometer:

- **Ionisation:**
 - o Electrospray $\rightarrow$ Sample dissolved in a highly volatile solvent and injected through a hypodermic needle that is attached to the positive terminal of a high voltage power supply. The particles gain H^+ and the solvent evaporates leaving gaseous positive ions.
 - o Electron impact $\rightarrow$ High energy electrons are fired out of an electron gun which knocks out electrons from atoms.
- **Acceleration:** A negatively charged plate attracts electrons towards it and all ions are accelerated to the same kinetic energy.
- **Flight tube:** The ions drift through this region with the same speed as they left the electric field with
- **Detection:** Ions hit a plate and gain an electron causing a current to flow. The size of this current can be measured to determine the abundance of each of the isotopes.

Mass spectrometry can be used to identify elements in a sample.



Figure 2: Diagram showing how a TOF mass spectrometer works.

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