

PIT100 Solutions to AIEEE-2007-CHEMISTRY Paper Code (O) 1

81. The energies of activation for forward and reverse reactions for $A_2 + B_2 \rightleftharpoons 2AB$ are 100 kJ mol^{-1} and 200 kJ mol^{-1} respectively. The presence of catalyst lowers the activation energy of both (forward and reverse) reactions by 100 kJ mol^{-1} . The enthalpy change of the reaction ($A_2 + B_2 \longrightarrow 2AB$) in the presence of catalyst will be (in kJ mol^{-1})

- (1) 300 (2) 120
(3) 280 (4) 20

Ans.
Sol.



$$\begin{aligned}\text{Sol. } \Delta H_{\text{reaction}} &= E_a - E_{ab} \\ &= 100 - 200 = -100\end{aligned}$$

Hence, (4) is correct.

82. The cell, $Zn(s) | Zn^{2+}(1M) || Cu^{2+}(1M) | Cu(s) | E_{\text{cell}}^{\circ} = 1.10V$, was allowed to be completely discharged at 298

- K. The relative concentration of Zn^{2+} to Cu^{2+} $\left[\left(\frac{Zn^{2+}}{Cu^{2+}} \right) \right]$ is
(1) antilog (24.08) (2) 37.3
(3) $10^{12.5}$ (4) 9.85×10^4

Ans.

$$\text{Sol. } E_{\text{cell}} = E_{\text{cell}}^{\circ} - \frac{0.0591}{n} \log Q$$

$$\text{where } Q = \left[\frac{Zn^{2+}}{Cu^{2+}} \right]$$

For complete discharge $E_{\text{cell}} = 0$

$$\text{So } E_{\text{cell}}^{\circ} = \frac{0.0591}{2} \log \left[\frac{Zn^{2+}}{Cu^{2+}} \right]$$

$$\Rightarrow \left[\frac{Zn^{2+}}{Cu^{2+}} \right] = 10^{12.5}$$

Hence, (3) is correct.

83. The pK_a of a weak acid (HA) is 4.5. The pOH of an aqueous buffered solution of HA in which 50% of the acid is ionized is

- (1) 4.5 (2) 2.5
(3) 9.5 (4) 7.0

Ans.

Sol. For buffer solution

$$pH = pK_a + \log \left[\frac{\text{Salt}}{\text{Acid}} \right]$$

$$= 4.5 + \log \left[\frac{\text{Salt}}{\text{Acid}} \right]$$

as HA is 50% ionized so $[\text{Salt}] = [\text{Acid}]$

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