

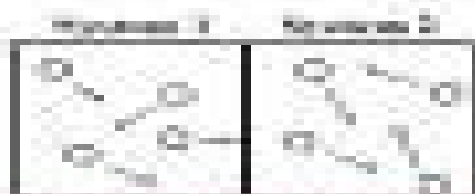
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The first curve of particles with maximum mass and average velocity contained all of the particles at 100 J . The thermally unstable mixture equal that the average velocity of all of the particles is 100 J (the average of the velocities of the particles is 100 J).

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, is the average kinetic temperature more System 1 or System 2? Justify your answer.

System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. The kinetic energy of the particles in the system is the same, but the particles in System 1 are moving faster. The particles in System 2 are moving slower. The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are moving faster than the particles in System 2.

- b. An initial amount of System 1 energy for System 2. Justify your answer in terms of molecular collisions.

System 1 is at a higher temperature than System 2. The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are moving faster than the particles in System 2.

Answer Key Unit 6 Thermochemistry

Arshad Iqbal



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