

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Procedure:

Calculations and Results:

1. Calculations:

(a)

$$\text{moles of sodium chloride} = 0.5 \text{ mol/L} \times 0.025 \text{ L} = 0.0125 \text{ mol}$$

(b)

$$\text{moles of calcium chloride} = 0.0125 \text{ mol} \times \frac{1}{2} = 0.00625 \text{ mol}$$

(c)

$$\text{mass of calcium chloride produced} = 0.00625 \text{ mol} \times 147.02 \text{ g/mol} = 0.919 \text{ g}$$

(d)

$$\text{mass of product as found (produced)} = 0.919 \text{ g} \times \frac{1}{1.00} = 0.919 \text{ g}$$

(e) Chemical equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

(f) Reaction:

Since: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$ This reaction is:

Na_2CO_3 and CaCl_2 are the reactants.

$$\text{moles of } \text{Na}_2\text{CO}_3 \text{ found in the } 0.0125 \text{ mol} = 0.0125 \text{ mol} \times \frac{1}{1} = 0.0125 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

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