

PhET Isotopes and Atomic Mass

Name Madison Fields

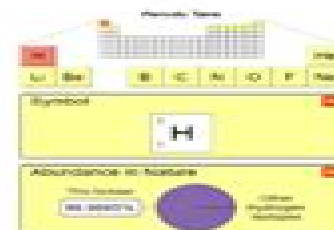
Date 10-2-2022

Use the link <https://phet.colorado.edu/en/simulation/isotopes-and-atomic-mass> to go to the PhET website for the simulation Isotopes and Atomic Mass.

Begin with the **Isotopes and Atomic Mass**.

Click on the + symbols next to the bar with Symbol and Abundance in Nature to display the data.

You can build isotopes of different atoms by clicking on the symbol and adding or subtracting neutrons using the bowl of neutrons next to the scale. You can also click on the scale to change to Atomic Mass (amu).



Protons: 5
Neutrons: 5
Electrons: 5

Let's practice with Boron.

1) What are the symbol, atomic mass and abundance in nature of the isotope Boron - 11?

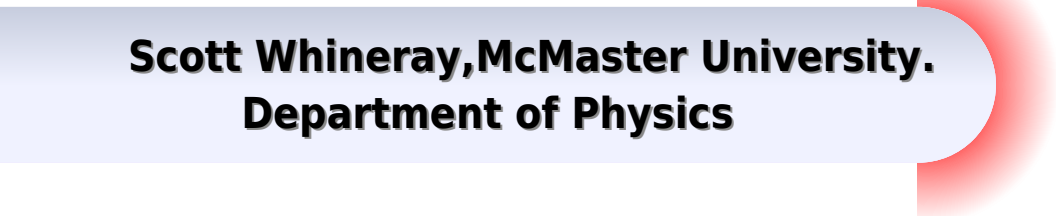


Now add one neutron to make Boron -12. Notice that the atom shakes. This indicates that the isotope is unstable. Unstable isotopes undergo radioactive decay. You will need to note this for other elements in the table below.

Symbol	$^{13}_6\text{C}$	$^{14}_7\text{N}$	$^{18}_8\text{O}$	$^{10}_4\text{Be}$	^4_2He	$^{19}_9\text{F}$
Name	Carbon-13		oxygen-18	Beryllium-10	helium-4	
Protons	6	7	8	4	2	9
Neutrons	7	6	10	6	2	10
Electrons	6	7	8	4	2	9
Atomic Mass	13.00335	13	17.99916	10.01353	4.00260	18.99840
Abundance	1.07%	0	0.205%	trace	99.9999	100%
Stable OR Unstable	stable	unstable	stable	unstable	stable	stable

Answers Isotopes And Atomic Mass Phet Sheet

**Scott Whineray, McMaster University.
Department of Physics**



Answers Isotopes And Atomic Mass Phet Sheet:

Preparing for Chemistry Teaching Festo Kayima, 2025-08-13 This textbook is a comprehensive chemistry didactics resource for chemistry teacher educators chemistry teachers and trainees It provides research grounded and practical based pedagogical experiences examples and frameworks for chemistry teachers as well as a foundation for planning and implementing productive chemistry lessons The book provides a conceptual and practical roadmap illuminating which didactic knowledge elements are relevant for becoming a chemistry teacher The book starts off with a pedagogically laden however experience based justification for the relevance of chemistry didactics and then progressively breaks down the different knowledge elements that form a complete set of the didactic knowledge and skill elements a teacher needs for teaching Concrete examples are provided to allow the reader to operationalize the ideas and concepts presented in the book The structure of the chapters enables the reader to engage progressively and actively with its contents and provided examples allowing a deep understanding of the diverse links between the presented topics forming a complete set of the didactic knowledge and skills relevant for successful chemistry teaching *Nuclear Science Abstracts*, 1965 **Structure of the Atom, Relative Atomic Mass & Isotopes**, 2006 *Table of Isotopes* Richard B. Firestone, 1999-08-03 Nuclear structure and decay data for thousands of isotopes and isomers a new update of the definitive reference The 1999 Table of Isotopes booklet set features Nuclear structure and radioactive decay data for approximately 3 700 isotopes and isomers an increase of more than 100 since the 1998 Update Up to date mass chain information with more than 10% revised material The latest versions as of December 1998 of the Evaluated Nuclear Structure Data File ENSDF and Nuclear Science Reference NSR file Additional data from several evaluation sources including The Table of Superdeformed Nuclear Bands and Fission Isomers Updated isotope summary table as well as energy ordered gamma ray and alpha particle tables now included on the software available from the book's website at www.wiley-vch.de/books/info/0_471_35633_6 Updated appendices for elemental data nuclear charts and gamma ray energy standards Adoption of the 1997 IUPAC recommended heavy element names More than 25 000 references Convenient links to additional atomic mass nuclear astrophysics rates spontaneous fission thermal neutron capture and more Plus the Isotope Explorer 2 22 software lets you search the entire database by level scheme drawings annotated tables data plots nuclear structure charts and keywords as well as download the latest data directly from the Table of Isotopes Web site The 1999 Update booklet features a new comprehensive isotope nuclear structure table In addition it provides clear step by step instructions on navigating the Table of Isotopes electronic content and accessing its Web site System Requirements PC Macintosh or UNIX systems with double speed and sufficient RAM to run Adobe TM Acrobat or see Adobe Acrobat Reader information in the book for specific system requirements Included Adobe Acrobat Reader or 3 02 for Windows or 95 98 and NT and Linux 1 2 13 or higher Acrobat Reader 3 01 for Windows 3 1 Isotope Explorer 2 22 for Windows 95 98 and NT *Atomic Mass Determinations for Some Isotopes of Ti, W*

and Hg Kenneth Stephen Kozier, 2012 *Atomic Mass Differences for Some Rare Earth Isotopes Near N* Bishop, Roy Lovitt, 1969 **Atomic Mass Determinations for Some Isotopes of Ti, W and Hg** Kozier, Kenneth Stephen, 1977 **Precise Atomic Mass Differences Amongst Isotopes of Hg, Tl, Pb, and Bi** Vladimir Peter Derenchuk, 1980 **Atomic Mass Differences for Some Rare Earth Isotopes Near N** Roy Lovitt Bishop, 2012 **Table of Isotopes** Richard B. Firestone, 1996 **Precise Atomic Mass Differences Involving Isotopes in the Heavy Mass Region Between Tungsten and Bismuth** William McLatchie, McMaster University. Department of Physics, 1966 **Precise Atomic Mass Differences Involving Isotopes of Dysprosium, Holmium, Erbium, Thulium, Ytterbium, Lutecium and Hafnium** Scott Whineray, McMaster University. Department of Physics, 1966 Radioactive Elements in the Standard Atomic Weights Table, 2011 In the 1949 Report of the Atomic Weights Commission a series of new elements were added to the Atomic Weights Table Since these elements had been produced in the laboratory and were not discovered in nature the atomic weight value of these artificial products would depend upon the production method Since atomic weight is a property of an element as it occurs in nature it would be incorrect to assign an atomic weight value to that element As a result of that discussion the Commission decided to provide only the mass number of the most stable or longest lived known isotope as the number to be associated with these entries in the Atomic Weights Table As a function of time the mass number associated with various elements has changed as longer lived isotopes of a particular element has been found in nature or as improved half life values of an element's isotopes might cause a shift in the longest lived isotope from one mass to another In the 1957 Report of the Atomic Weights Commission it was decided to discontinue the listing of the mass number in the Atomic Weights Table on the grounds that the kind of information supplied by the mass number is inconsistent with the primary purpose of the Table i.e. to provide accurate values of these constants for use in various chemical calculations In addition to the Table of Atomic Weights the Commission included an auxiliary Table of Radioactive Elements for the first time where the entry would be the isotope of that element which was the most stable i.e. the one with the longest known half life In their 1973 Report the Commission noted that the users of the main Table of Atomic Weights were dissatisfied with the omission of values for some elements in that Table and it was decided to reintroduce the mass number for the radioactive elements into the main Table In their 1983 Report the Commission decided that radioactive elements were considered to lack a characteristic terrestrial isotopic composition from which an atomic weight value could be calculated to five or more figure accuracy without prior knowledge of the sample involved These elements were again listed in the Atomic Weights Table with no further information i.e. with no mass number or atomic weight value For the elements which have no stable characteristic terrestrial isotopic composition the data on the half lives and the relative atomic masses for the nuclides of interest for those elements have been evaluated The values of the half lives with their uncertainties are listed in the table The uncertainties are given for the last digit quoted of the half life and are given in parentheses A half life entry for the Table having a value and an uncertainty

of ^{73}As is listed in the half life column as 73 The criteria to include data in this Table is to be the same as it has been for over sixty years It is the same criteria which are used for all data that are evaluated for inclusion in the Standard Table of Atomic Weights If a report of data is published in a peer reviewed journal that data is evaluated and considered for inclusion in the appropriate table of the biennial report of the Atomic Weights Commission As better data becomes available in the future the information that is contained in either of the Tables of Standard Atomic Weights or in the Table of Radioactive Elements may be modified It should be noted that the appearance of any datum in the Table of the Radioactive Elements is merely for the purposes of calculating an atomic mass value for any sample of a radioactive material which might have a variety of isotopic compositions and it has no implication as to the priority for claiming discovery of a given element and is not intended to The atomic mass values have been taken primarily from the 2003 Atomic Mass Table Mass values for those radioisotopes that do not appear in the 2003 Atomic mass Table have been taken from preliminary data of the Atomic Mass Data Center Most of the quoted half lives

Precise Atomic Mass Differences Amongst Isotopes of Hg, Tl, Pb, and Bi Vladimir Peter Derenchuk,1980

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