

2009 AP® CHEMISTRY FREE-RESPONSE QUESTIONS

Answer Question 5 and Question 6. The Section II score weighting for these questions is 15 percent each.

Your responses to these questions will be graded on the basis of the accuracy and relevance of the information cited. Explanations should be clear and well organized. Examples and equations may be included in your responses where appropriate. Specific answers are preferable to broad, diffuse responses.

Reaction	Equation	ΔH_{298}°	ΔS_{298}°	ΔG_{298}°
X	$\text{C}(s) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}(g) + \text{H}_2(g)$	$+131 \text{ kJ mol}^{-1}$	$+134 \text{ J mol}^{-1} \text{ K}^{-1}$	$+91 \text{ kJ mol}^{-1}$
Y	$\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$	$+41 \text{ kJ mol}^{-1}$	$+42 \text{ J mol}^{-1} \text{ K}^{-1}$	$+29 \text{ kJ mol}^{-1}$
Z	$2 \text{ CO}(g) \rightleftharpoons \text{C}(s) + \text{CO}_2(g)$	?	?	?

5. Answer the following questions using the information related to reactions X, Y, and Z in the table above.

- For reaction X, write the expression for the equilibrium constant, K_p .
- For reaction X, will the equilibrium constant, K_p , increase, decrease, or remain the same if the temperature rises above 298 K? Justify your answer.
- For reaction Y at 298 K, is the value of K_p greater than 1, less than 1, or equal to 1? Justify your answer.
- For reaction Y at 298 K, which is larger: the total bond energy of the reactants or the total bond energy of the products? Explain.
- Is the following statement true or false? Justify your answer.
 "On the basis of the data in the table, it can be predicted that reaction Y will occur more rapidly than reaction X will occur."
- Consider reaction Z at 298 K.
 - Is ΔS° for the reaction positive, negative, or zero? Justify your answer.
 - Determine the value of ΔH° for the reaction.
 - A sealed glass reaction vessel contains only $\text{CO}(g)$ and a small amount of $\text{C}(s)$. If a reaction occurs and the temperature is held constant at 298 K, will the pressure in the reaction vessel increase, decrease, or remain the same over time? Explain.

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Table of Contents Answers Ap Chemistry 2013 Response Questions

1. Understanding the eBook Answers Ap Chemistry 2013 Response Questions
 - The Rise of Digital Reading Answers Ap Chemistry 2013 Response Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Ap Chemistry 2013 Response Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Ap Chemistry 2013 Response Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Ap Chemistry 2013 Response Questions
 - Personalized Recommendations
 - Answers Ap Chemistry 2013 Response Questions User Reviews and Ratings

- Answers Ap Chemistry 2013 Response Questions and Bestseller Lists
- 5. Accessing Answers Ap Chemistry 2013 Response Questions Free and Paid eBooks
 - Answers Ap Chemistry 2013 Response Questions Public Domain eBooks
 - Answers Ap Chemistry 2013 Response Questions eBook Subscription Services
 - Answers Ap Chemistry 2013 Response Questions Budget-Friendly Options
- 6. Navigating Answers Ap Chemistry 2013 Response Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Ap Chemistry 2013 Response Questions Compatibility with Devices
 - Answers Ap Chemistry 2013 Response Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Ap Chemistry 2013 Response Questions
 - Highlighting and Note-Taking Answers Ap Chemistry 2013 Response Questions
 - Interactive Elements Answers Ap Chemistry 2013 Response Questions
- 8. Staying Engaged with Answers Ap Chemistry 2013 Response Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Ap Chemistry 2013 Response Questions
- 9. Balancing eBooks and Physical Books Answers Ap Chemistry 2013 Response Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Ap Chemistry 2013 Response Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Ap Chemistry 2013 Response Questions
 - Setting Reading Goals Answers Ap Chemistry 2013 Response Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Ap Chemistry 2013 Response Questions
 - Fact-Checking eBook Content of Answers Ap Chemistry 2013 Response Questions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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